

A Highly Effective Catalyst System for the Pd-Catalyzed Amination of Vinyl Bromides and Chlorides

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Supporting Information

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General Considerations

All reactions were performed under an atmosphere of argon in oven-dried glassware. Toluene was freshly distilled over sodium/benzophenone and stored over 4 Å molecular sieves under an argon atmosphere. ^1H and ^{13}C NMR spectra were recorded at 300 and 75.5 MHz, respectively. Thin layer chromatography (TLC) was performed using commercially prepared 60 mesh silica gel plates visualized with short-wavelength UV light (254 nm). For convenience, stock solutions of ligands (2 mM) were prepared in toluene and stored under argon. Celite[®] 521 (Aldrich) was used for the purification of enamines. Electron impact ionization experiments were performed on a Finnigan TSQ700 triple quadrupole mass spectrometer (Finnigan MAT, San Jose, CA) fitted with a Finnigan EI/CI ion source. Accurate mass measurements were performed using a double focusing Kratos MS-50 mass spectrometer (Kratos, NJ). The reported yields are isolated yields and are the average of two runs. All commercially available reagents were used as received. Ligands **1a**, **1b** and **1d** were received from Aldrich and used without purification, ligands **1c** and **1e** were prepared according to our previous procedures.¹ All compounds described in Tables 1-4 and Scheme 1 are known in the literature (unless stated otherwise) and were characterized by comparing their ^1H and ^{13}C NMR to the previously reported data. In all cases, the comparisons were very favorable. Compound **5e** is new and it was characterized by ^1H , ^{13}C , mass (EI) and HRMS analysis. Compound **8b** is known, but since data are not available it was characterized in the same manner as **5e**.

Experimental Procedures

General Procedure for the Synthesis of Enamines from Vinyl Bromides. An oven-dried Schlenk flask equipped with a magnetic stirring bar was charged with NaO-*t*-Bu (1.4 equiv.) and Pd(OAc)₂ or Pd₂(dba)₃ (0.25 mol %) inside a nitrogen-filled glove box. The flask was capped with a rubber septum, and then it was removed from the glove box. Alkenyl bromide (3.0 mmol), amine (3.0 mmol), ligand (0.5 mol %) and toluene (5 mL)

were then successively added. The Schlenk flask was degassed and refilled with argon. Then the flask was placed in an 80 °C oil bath, and the reaction mixture was stirred until the starting material had been completely consumed as judged by TLC. After completion of the reaction, the mixture was cooled to room temperature, diluted with 30 mL of hexanes and filtered through Celite[®]. Solvents were evaporated under reduced pressure to afford a residue that consisted of the essentially pure enamine (¹H and ¹³C NMR).

General Procedure for the Synthesis of Imines from Vinyl Bromides. An oven-dried Schlenk flask equipped with a magnetic stirring bar was charged with NaO-*t*-Bu (1.4 equiv.) and Pd₂(dba)₃ (0.25 mol %) inside a nitrogen-filled glove box. The flask was capped with a rubber septum, and then it was removed from the glove box. Alkenyl bromide (3.0 mmol), amine (3.0 mmol), ligand (0.5 mol %) and toluene (5 mL) were then successively added. The Schlenk flask was degassed and refilled with argon. Then the flask was placed in an 80 °C oil bath, and the reaction mixture was stirred until the starting material had been completely consumed as judged by TLC. After completion of the reaction, the mixture was cooled to room temperature, diluted with 30 mL of hexanes and filtered through Celite[®]. Solvents were evaporated under reduced pressure to afford a residue that consisted of the essentially pure enamine (¹H and ¹³C NMR).

Compound **7d** was purified as follows: After completion of the reaction, the mixture was cooled to room temperature, diluted with 30 mL of hexanes and filtered through Celite[®]. Solvents were evaporated, the obtained residue was separated by flash chromatography through neutral alumina (25 g) treated with Et₃N (5 g), and then eluted with 200 mL hexanes.

General Procedure for the Synthesis of Enamines from Vinyl Chloride 8. An oven-dried Schlenk flask equipped with a magnetic stirring bar was charged with NaO-*t*-Bu (1.4 equiv.) and Pd₂(dba)₃ (2.5 mol %) inside a nitrogen-filled glove box. The flask was capped with a rubber septum, and then it was removed from the glove box. Vinyl chloride **8** (3.0 mmol), amine (3.0 mmol), ligand (5 mol %) and toluene (5 mL) were then successively added. The Schlenk flask was degassed and refilled with argon. Then, the flask was placed in a 115 °C oil bath and the reaction mixture was stirred until the starting

material had been completely consumed as judged by TLC. After completion of the reaction, the mixture was cooled to room temperature, diluted with 30 mL of hexanes and filtered through Celite[®]. Solvents were evaporated under reduced pressure to afford a residue, which consisted of the essentially pure enamine (¹H and ¹³C NMR).

General Procedure for the Synthesis of 1*H*-Indole 11 by Palladium Catalyzed Cascade Reaction of Alkenyl Bromide 2 with *o*-Bromoaniline 10. An oven-dried Schlenk flask equipped with a magnetic stirring bar was charged with NaO-*t*-Bu (2.8 equiv.), *o*-bromoaniline **10** (3 mmol) and Pd₂(dba)₃ (4.0 mol %) inside a nitrogen-filled glove box. The flask was capped with a rubber septum, and removed from the glove box. α -Bromostyrene **2** (3.0 mmol), ligand **1d** (4 mol %) and toluene (5 mL) were then successively added. The Schlenk flask was degassed and refilled with argon. The flask was then placed in an 80 °C oil bath and the reaction mixture was stirred until **10** had been completely consumed as judged by TLC. The temperature was then raised to 120 °C and after completion of the reaction (16 h), the mixture was cooled to room temperature, diluted with 50 mL of hexanes and filtered through Celite[®]. Solvents were evaporated under reduced pressure. Purification by silica gel flash chromatography (hexanes/ethyl acetate, 10:1) afforded 1*H*-indole **11**.

(1) Su, W.; Urgaonkar, S.; McLaughlin, P. A.; Verkade, J. G. *J. Am. Chem. Soc.* **2004**, *126*, 16433.

References for known compounds

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N-(2-Phenylethyl)-piperidine (Table 3, **5b**). Tillack, A.; Trauthwein, H.; Hartung, C. G.; Eichberger, M.; Jansen, A.; Bellar, M. *Monatsh. Chemie*, **2000**, 131, 1327.

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N-(2-Phenylethyl)-diethylamine (Table 3, **5h**). Tillack, A.; Trauthwein, H.; Hartung, C. G.; Eichberger, M.; Jansen, A.; Bellar, M. *Monatsh. Chemie*, **2000**, 131, 1327.

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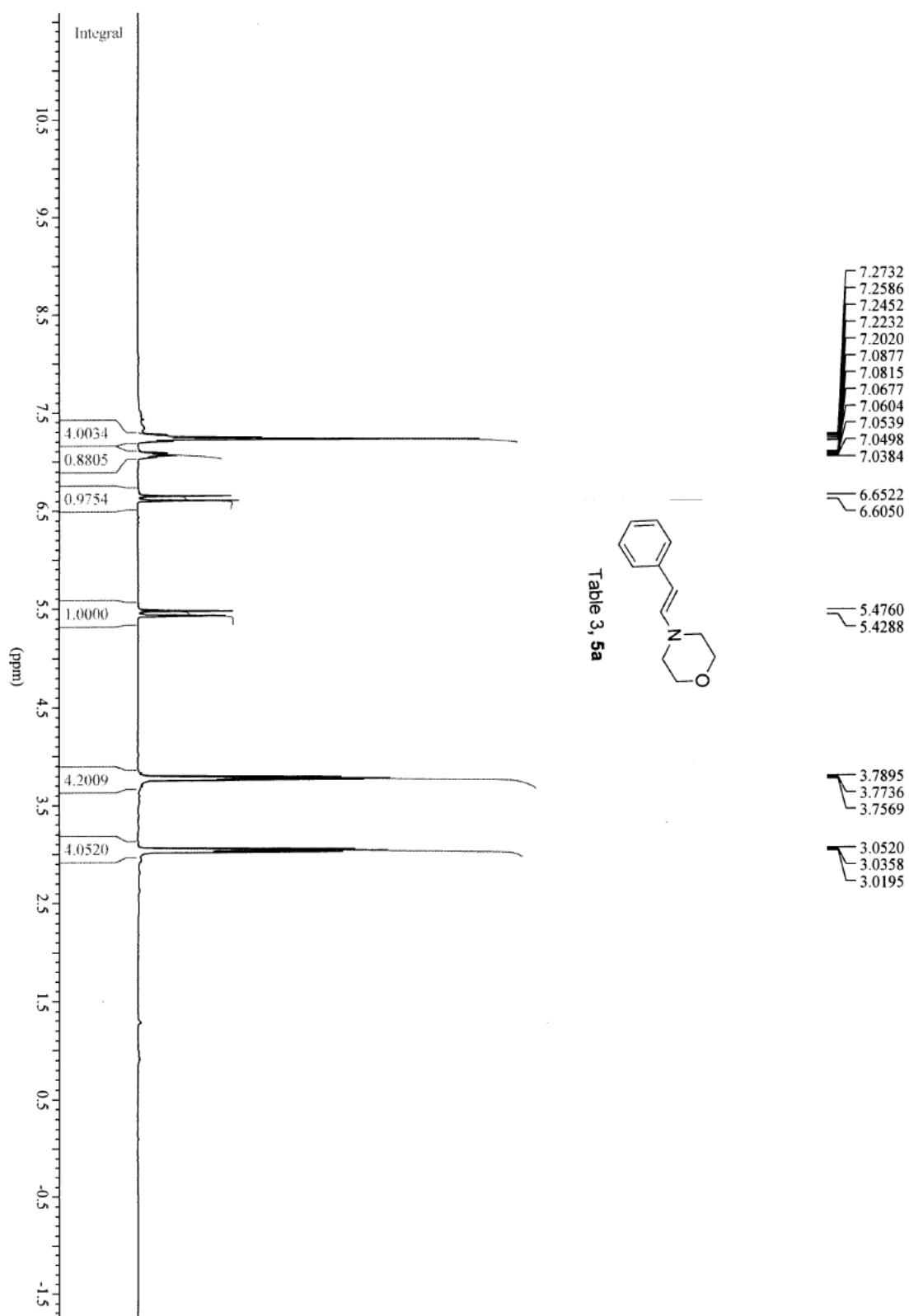
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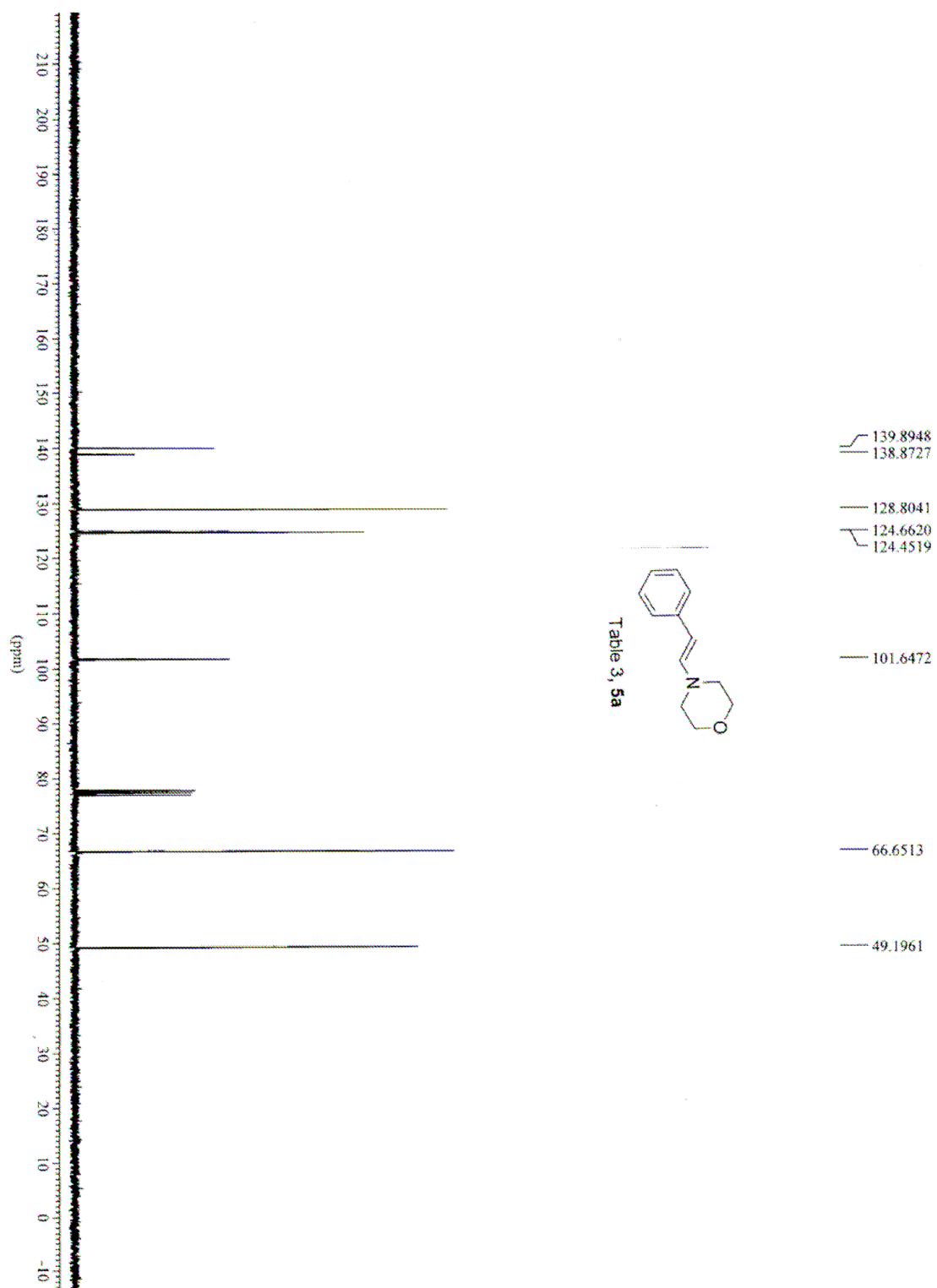
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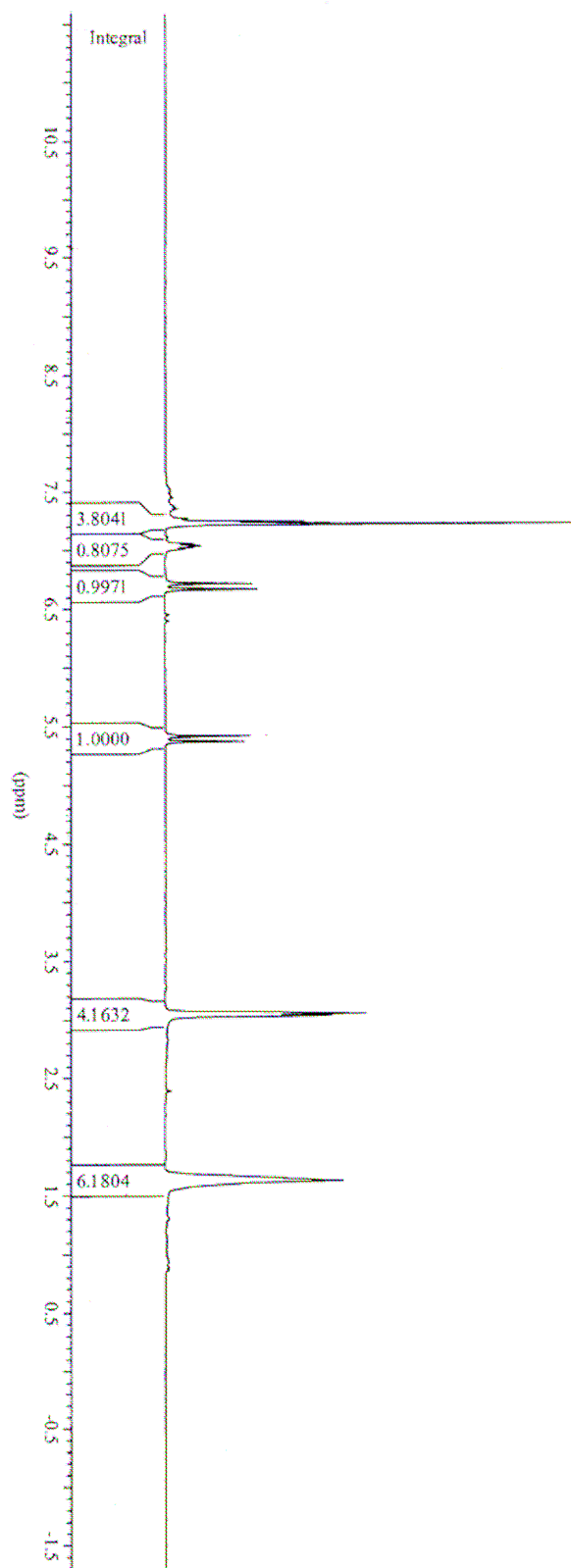
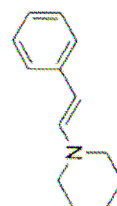


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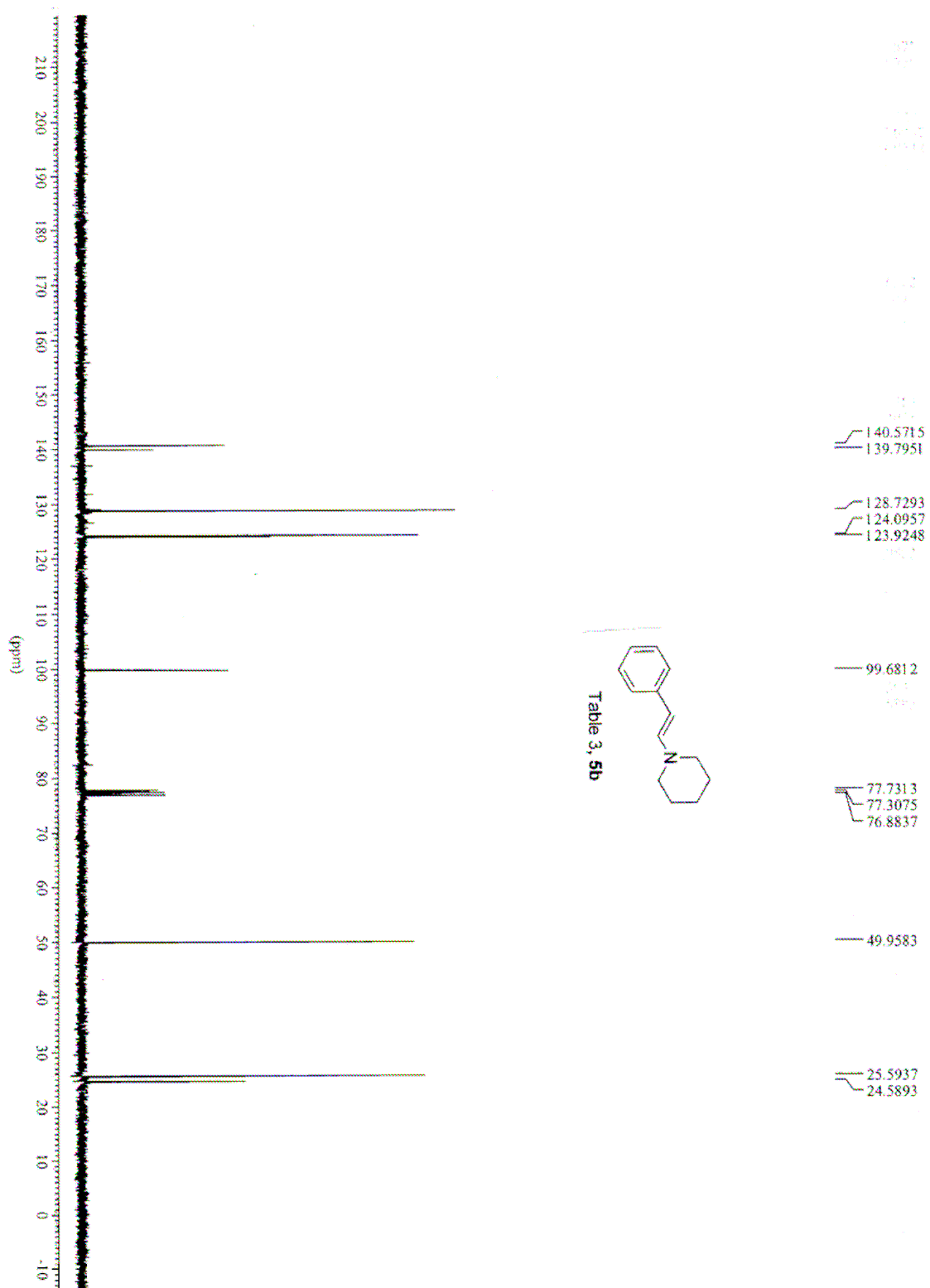


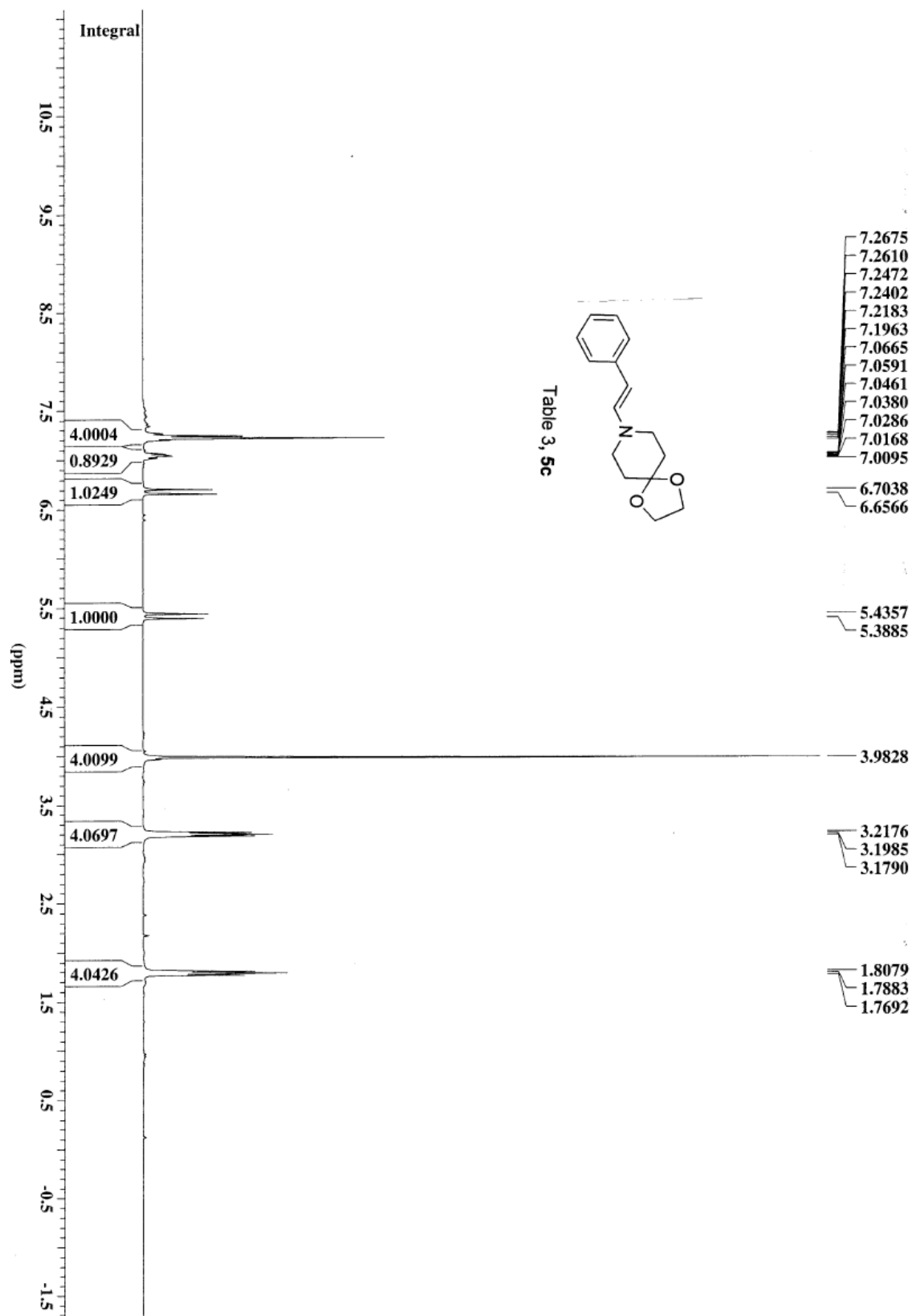
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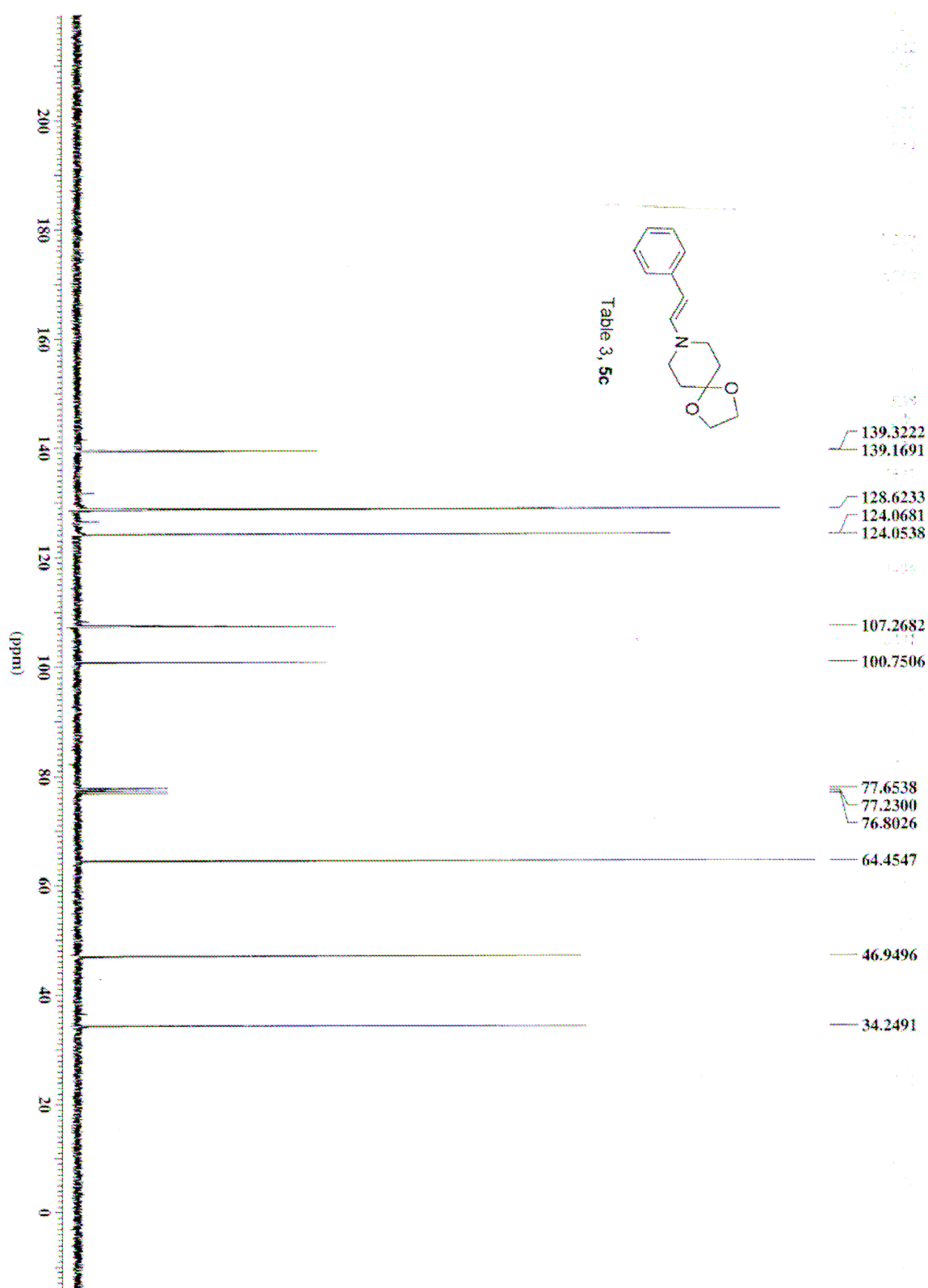
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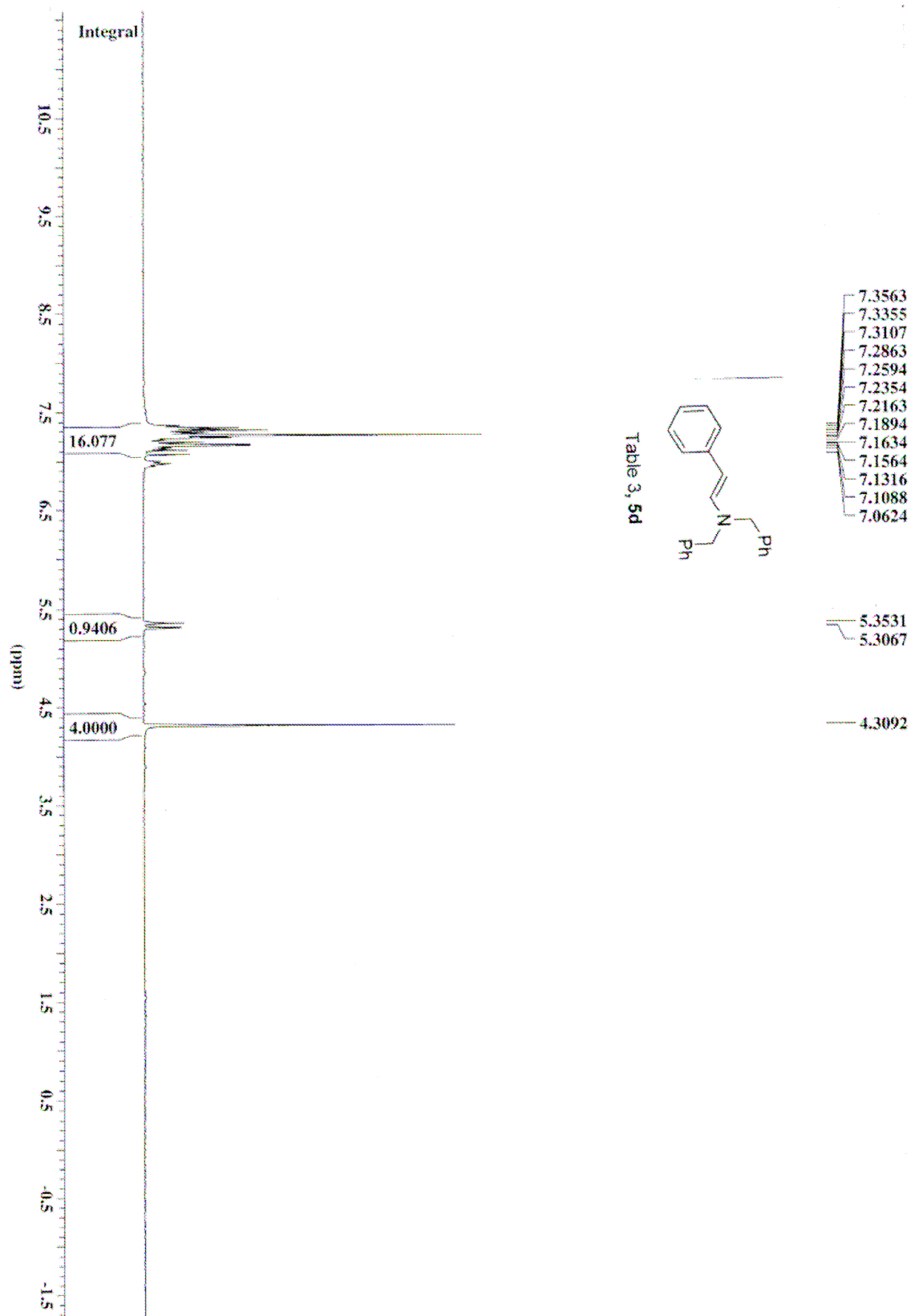
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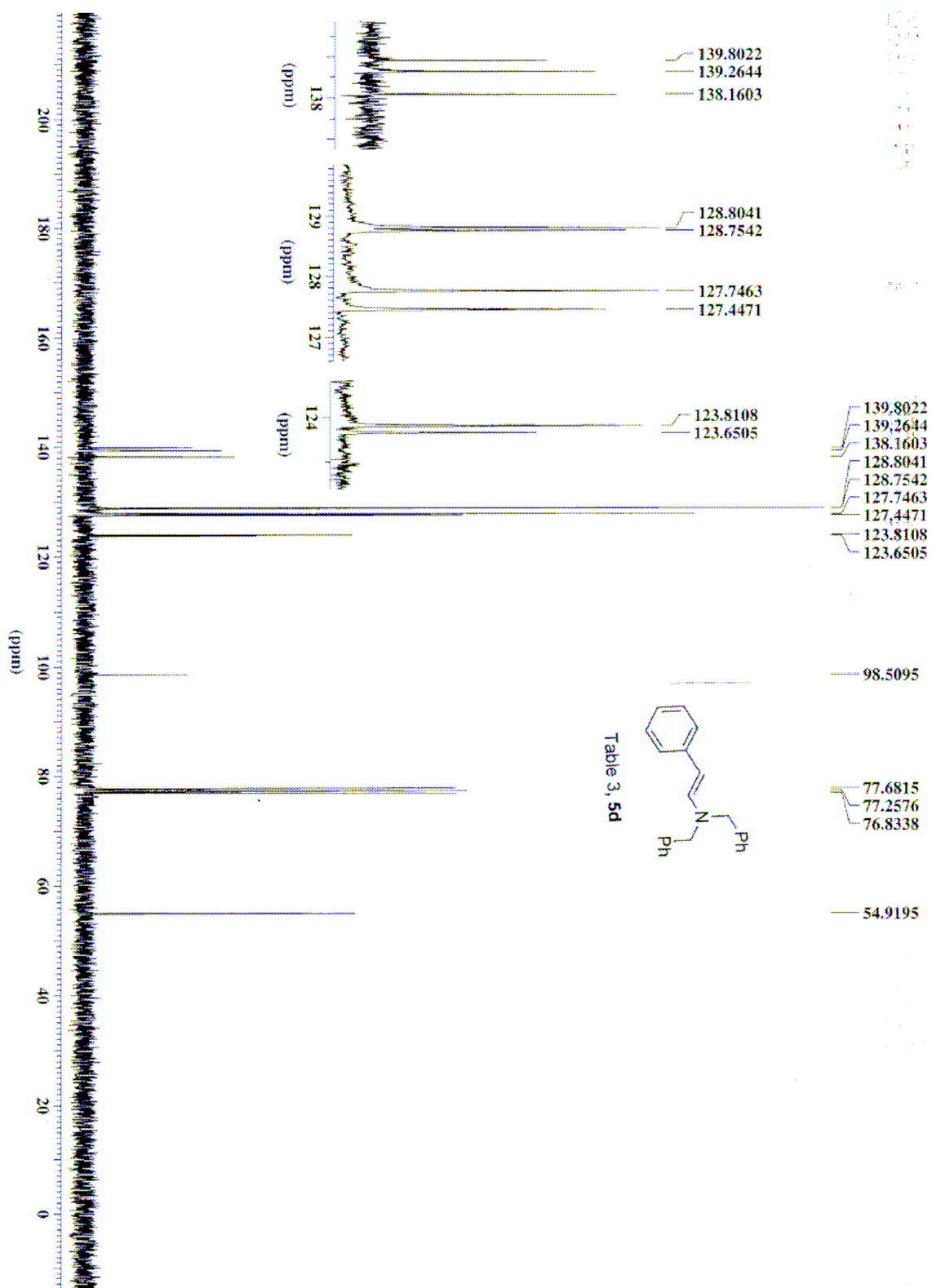
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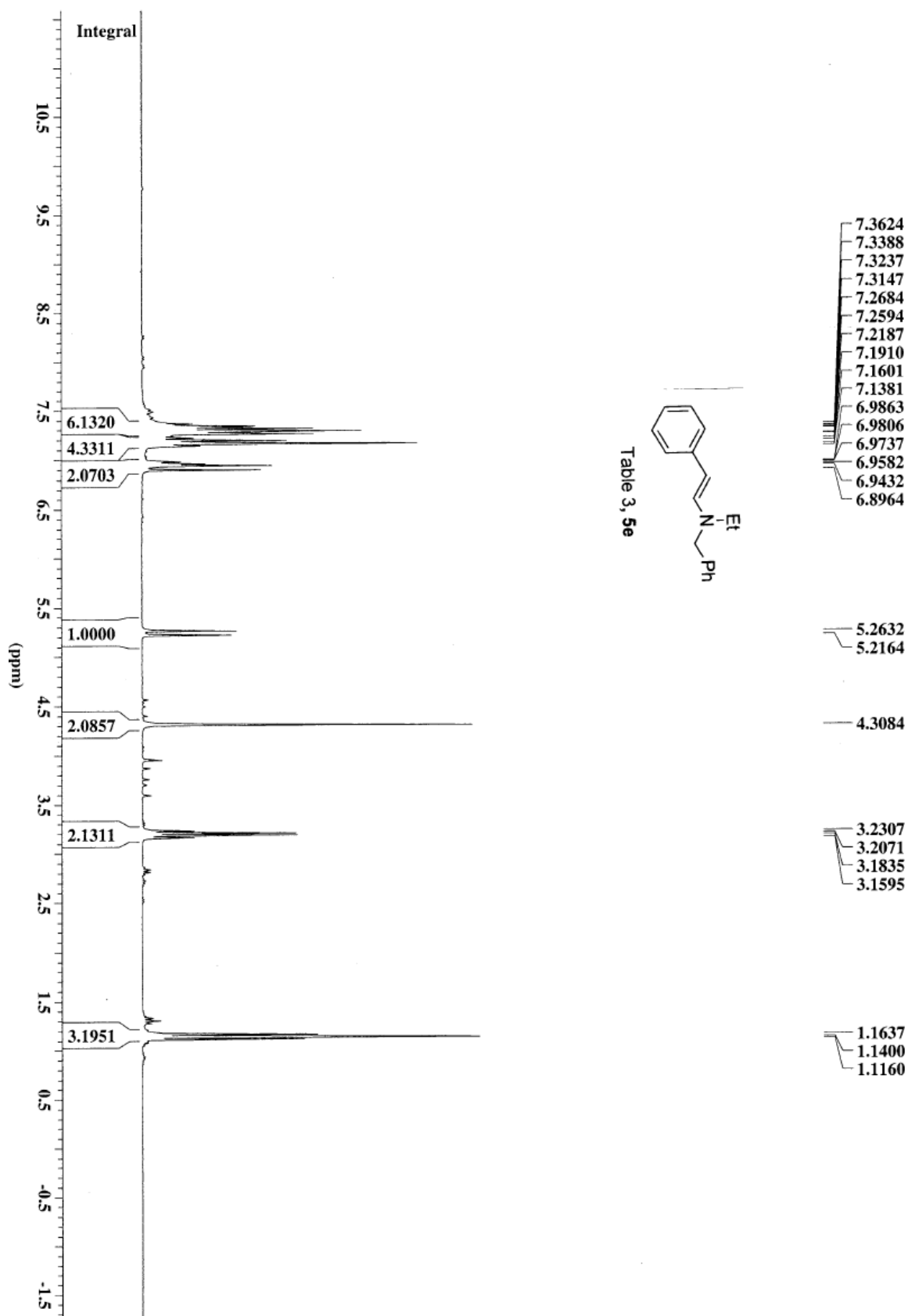


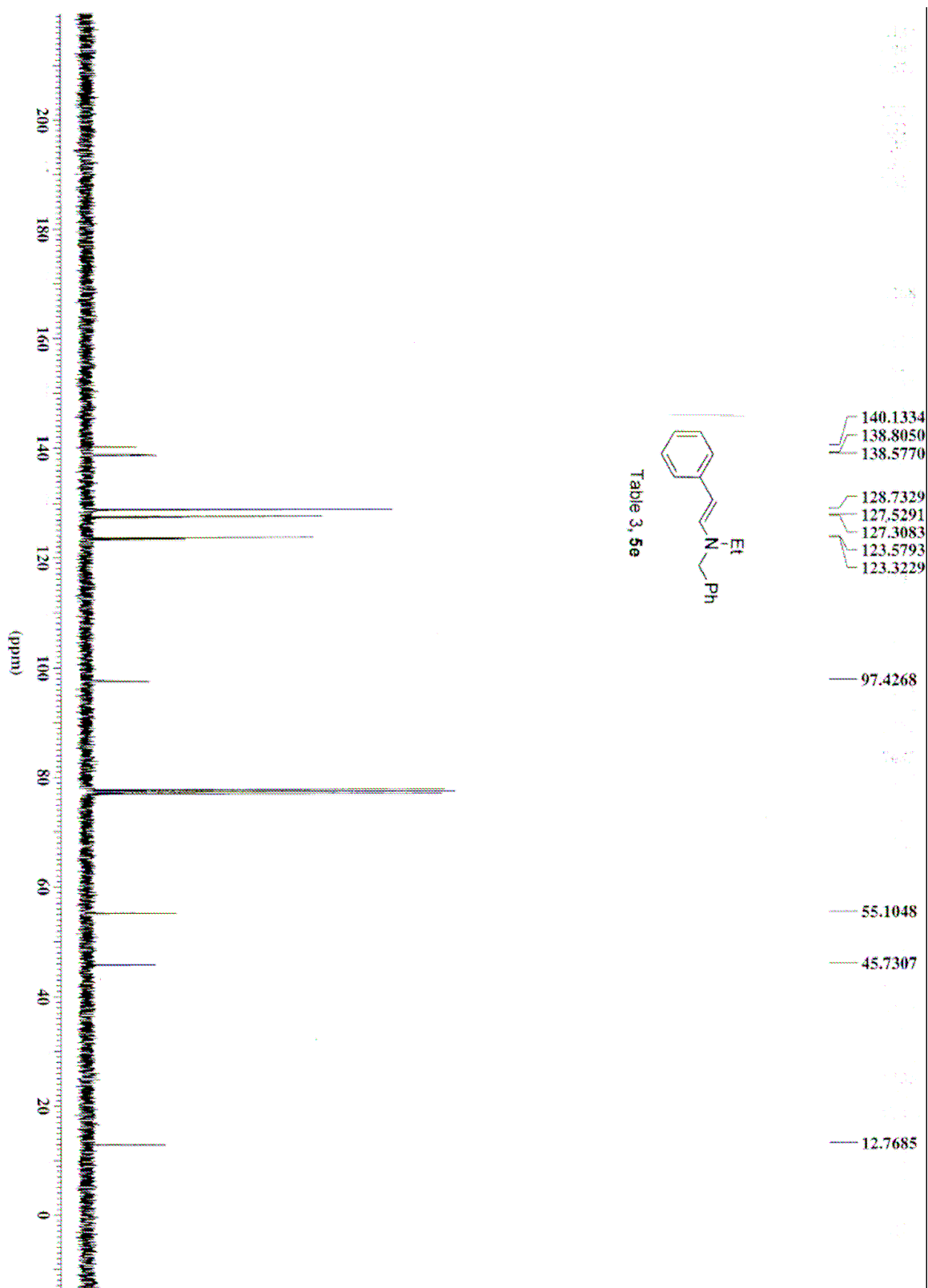






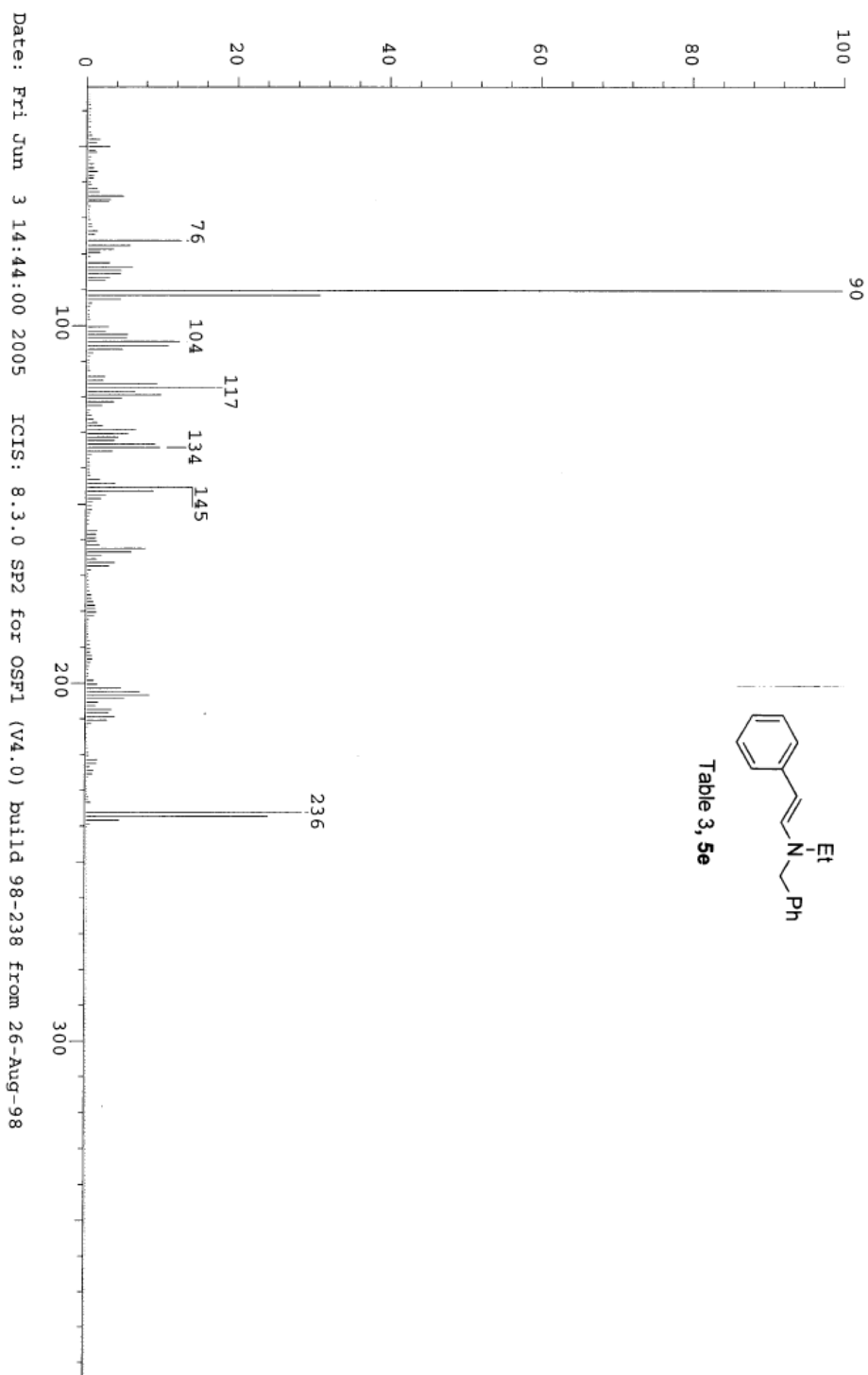
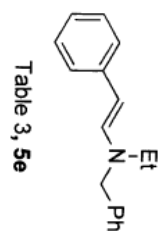






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699017					
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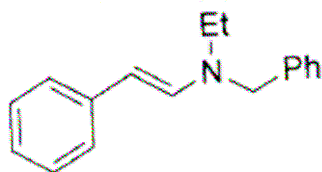


Table 3, 5e

Manual Peak Matching Report For Accurate Mass Determination

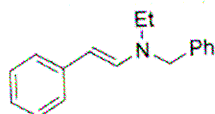


Table 3, 5e

Theoretical mass	Experimental mass	PFK matching mass	Deviation*
237.15175	237.15205	230.98562	1.3 ppm

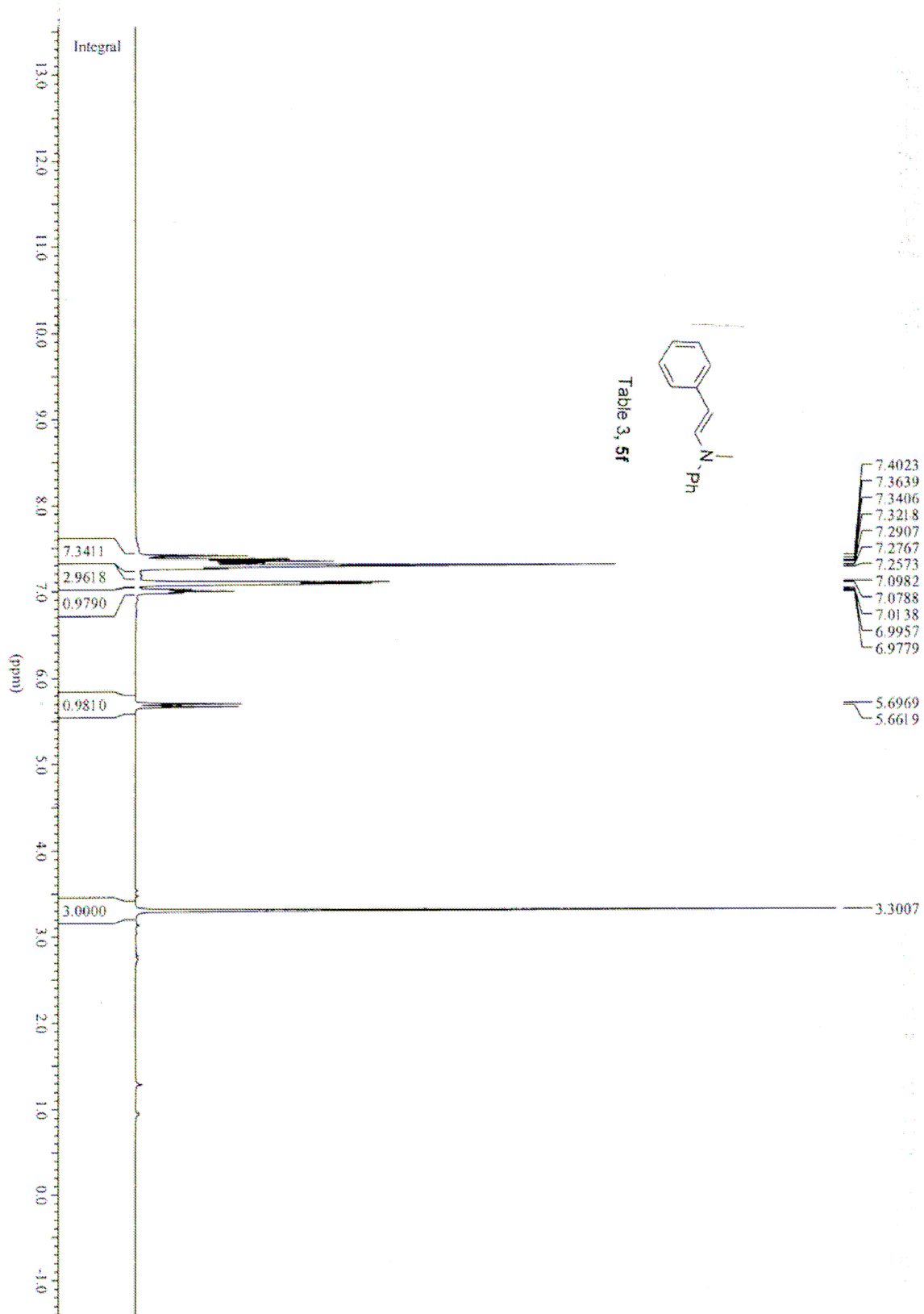
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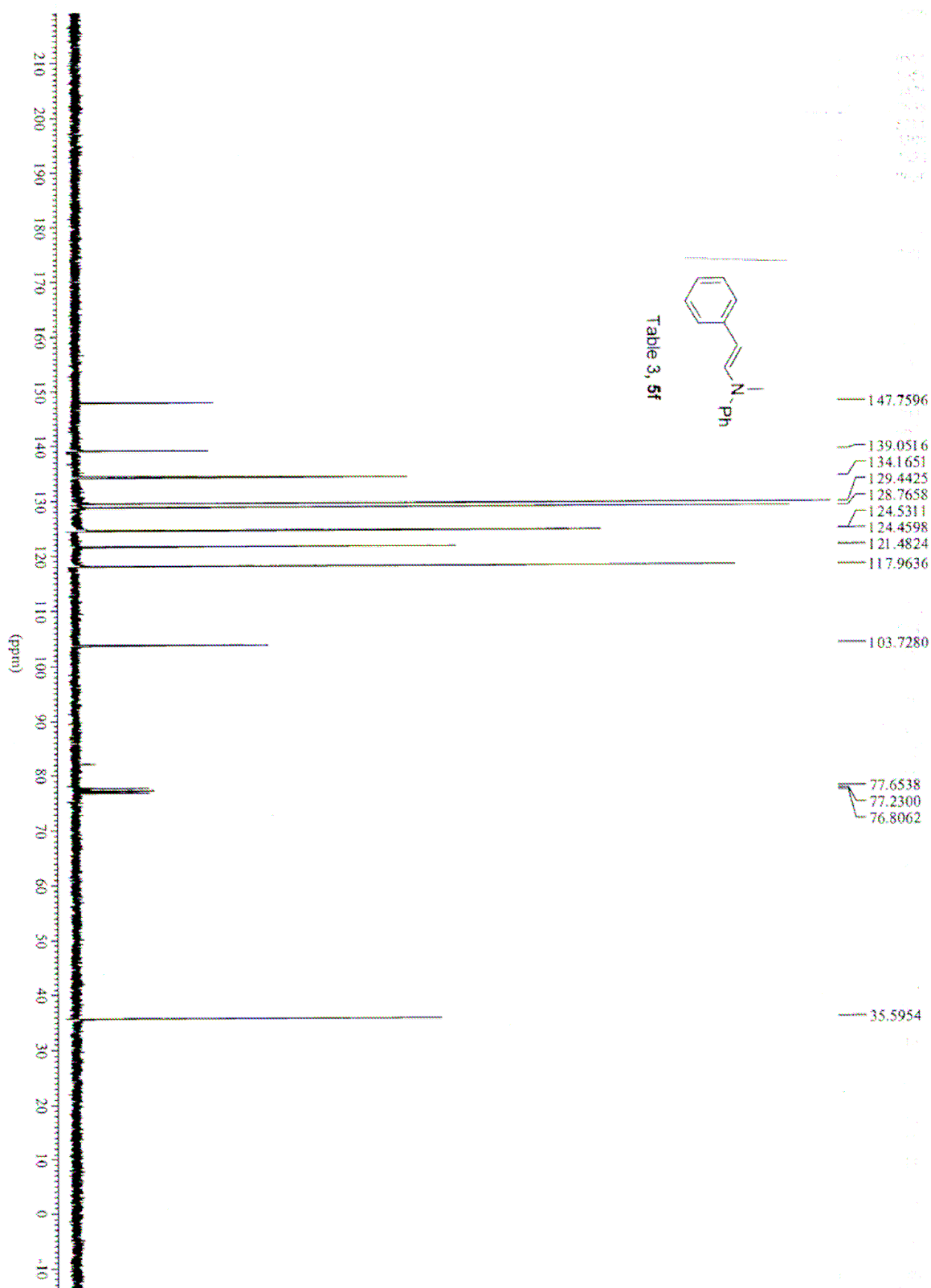
$$\text{deviation} = \frac{\text{experimental mass} - \text{theoretical mass}}{\text{nominal mass}}$$

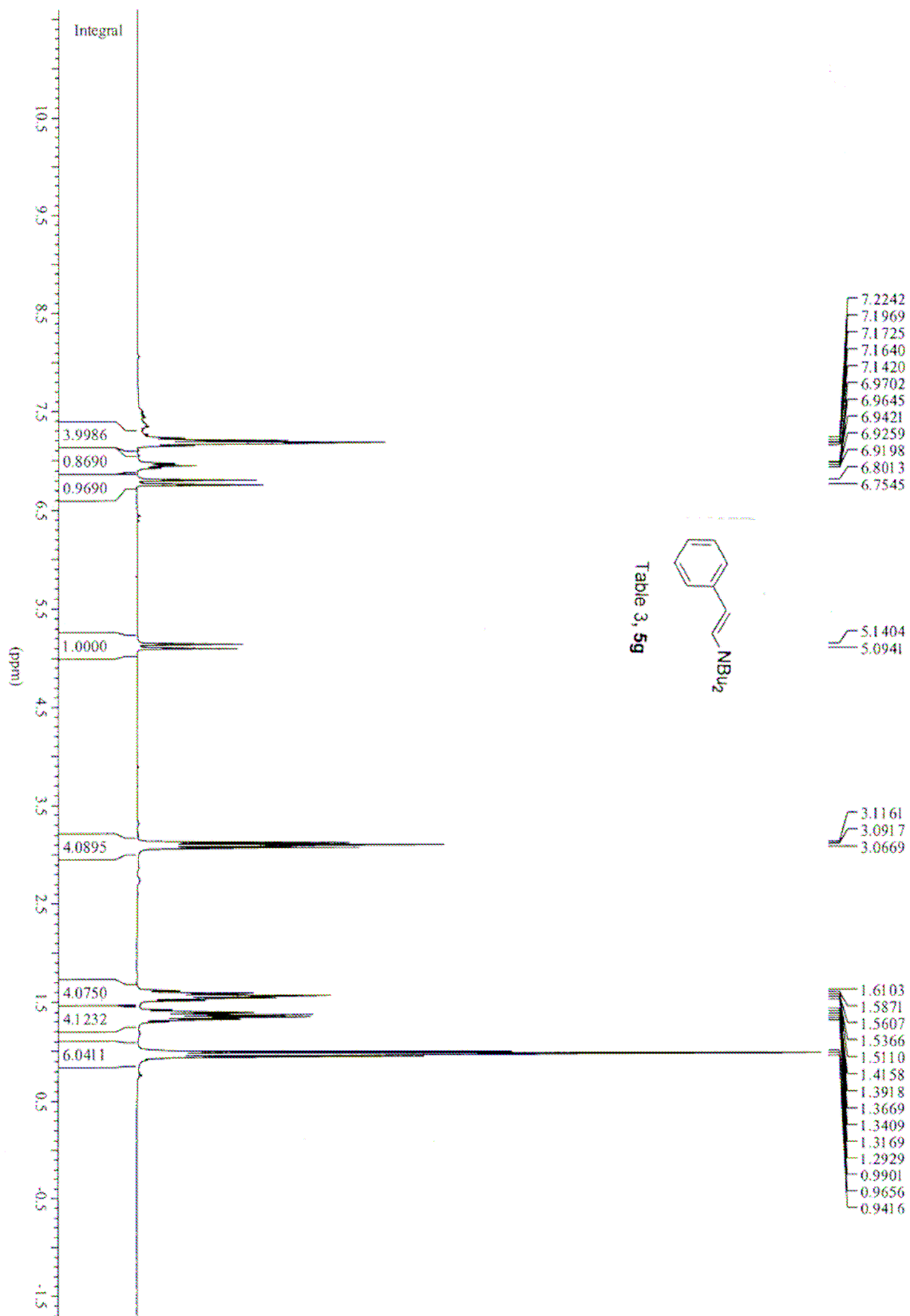
Where nominal mass takes in account only ^{12}C , ^1H , ^{16}O , ^{14}N etc...

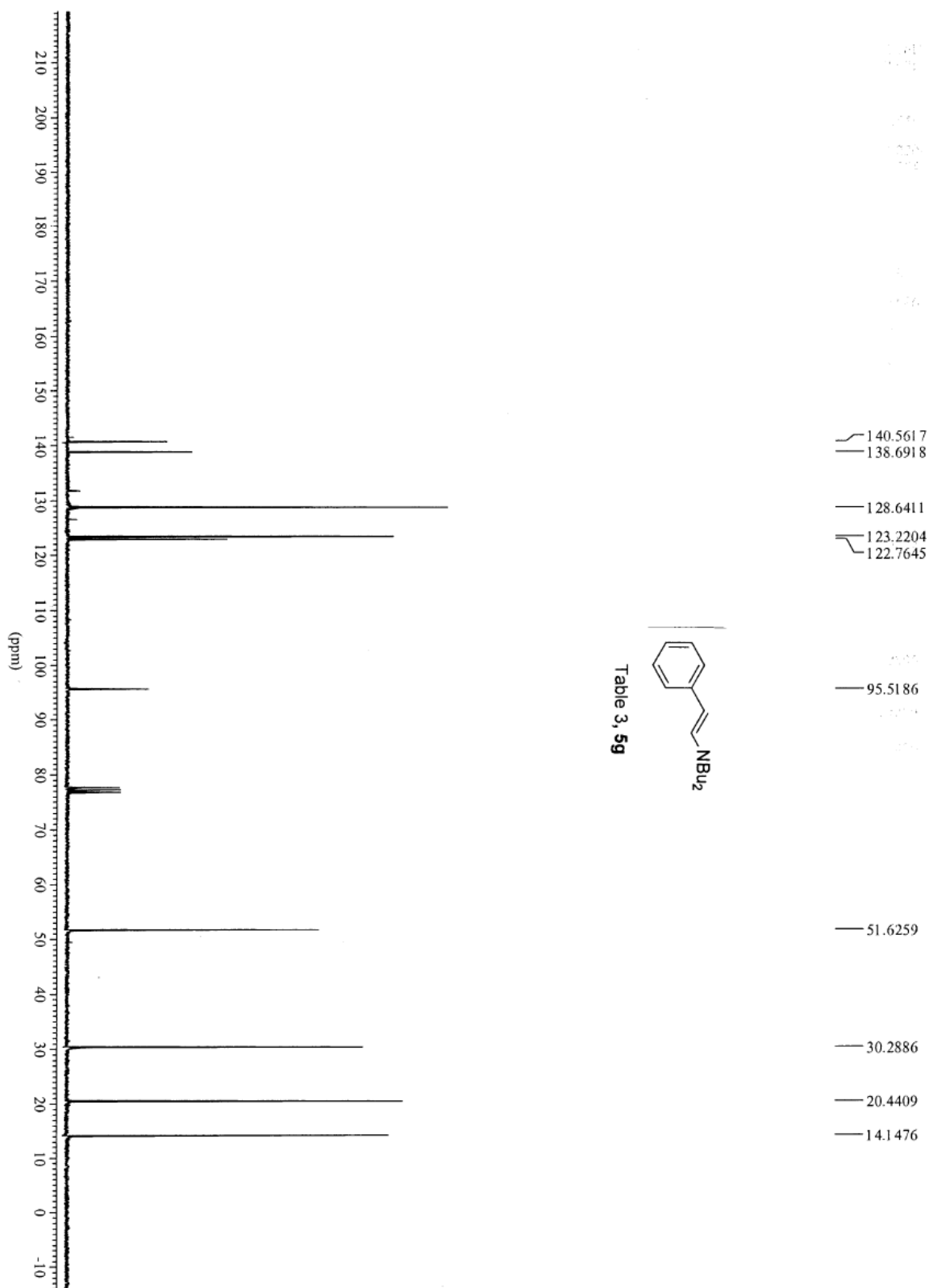
Theoretical mass correspond to the mass of the most abundant isotope peak

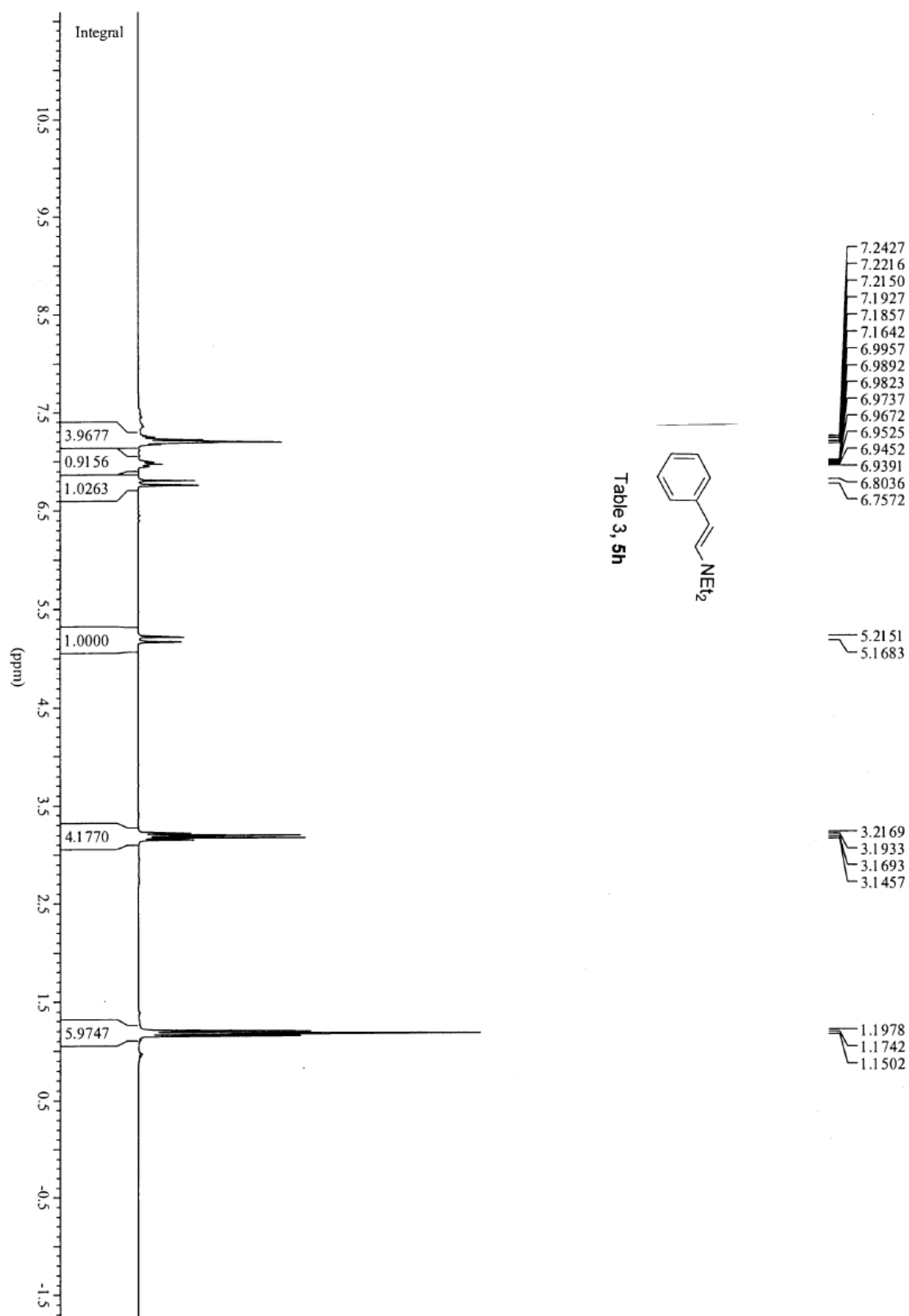
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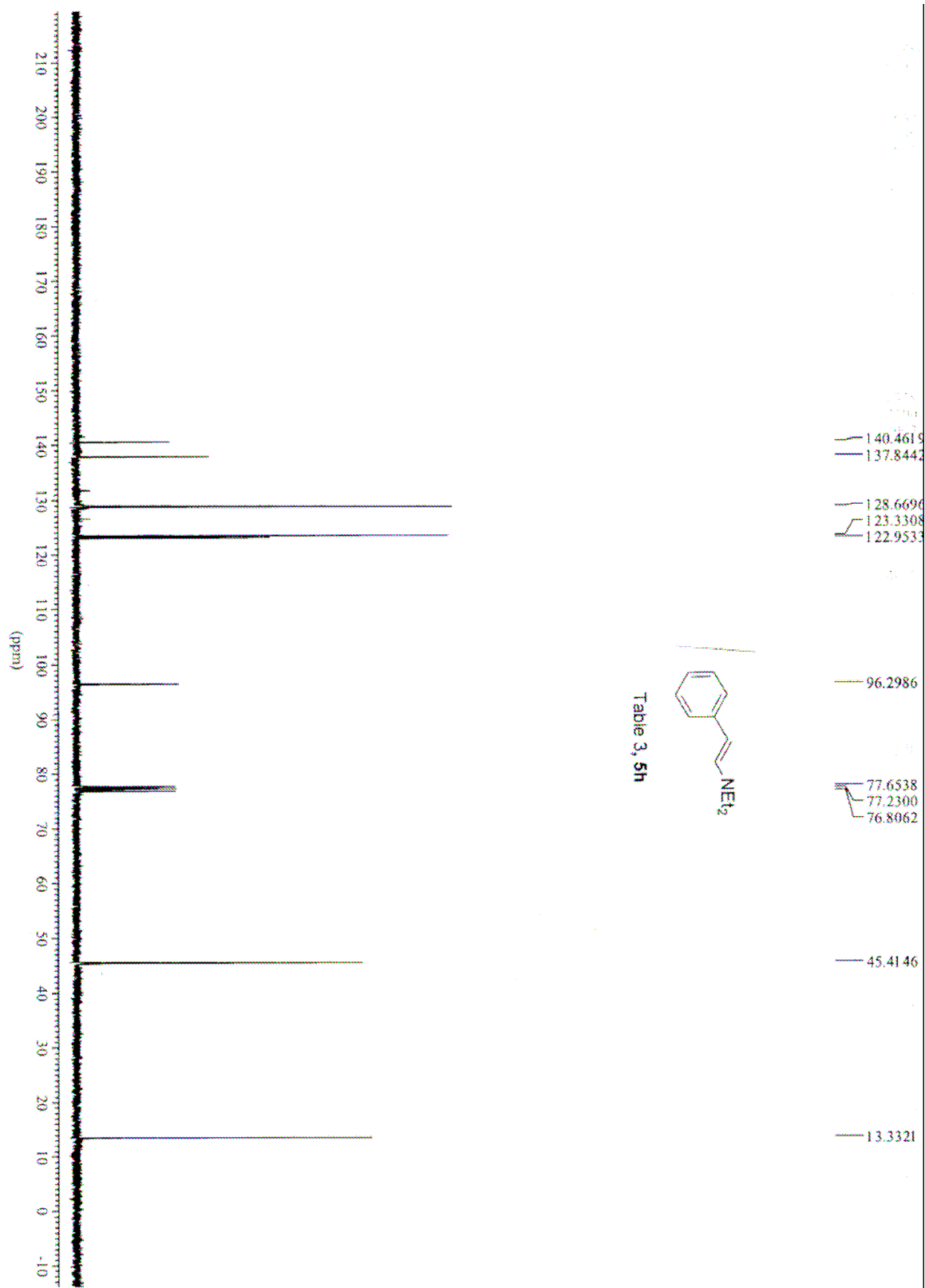


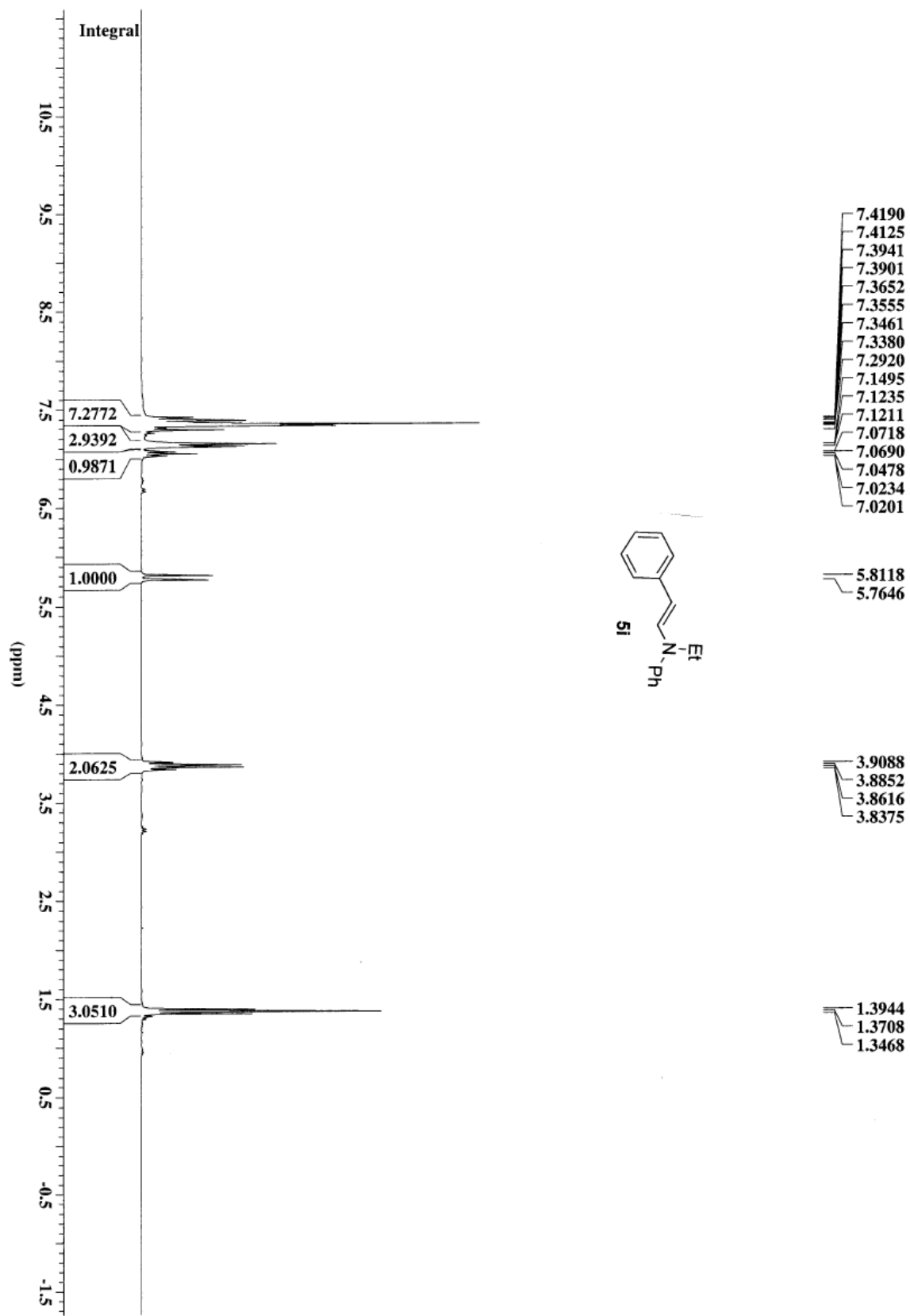


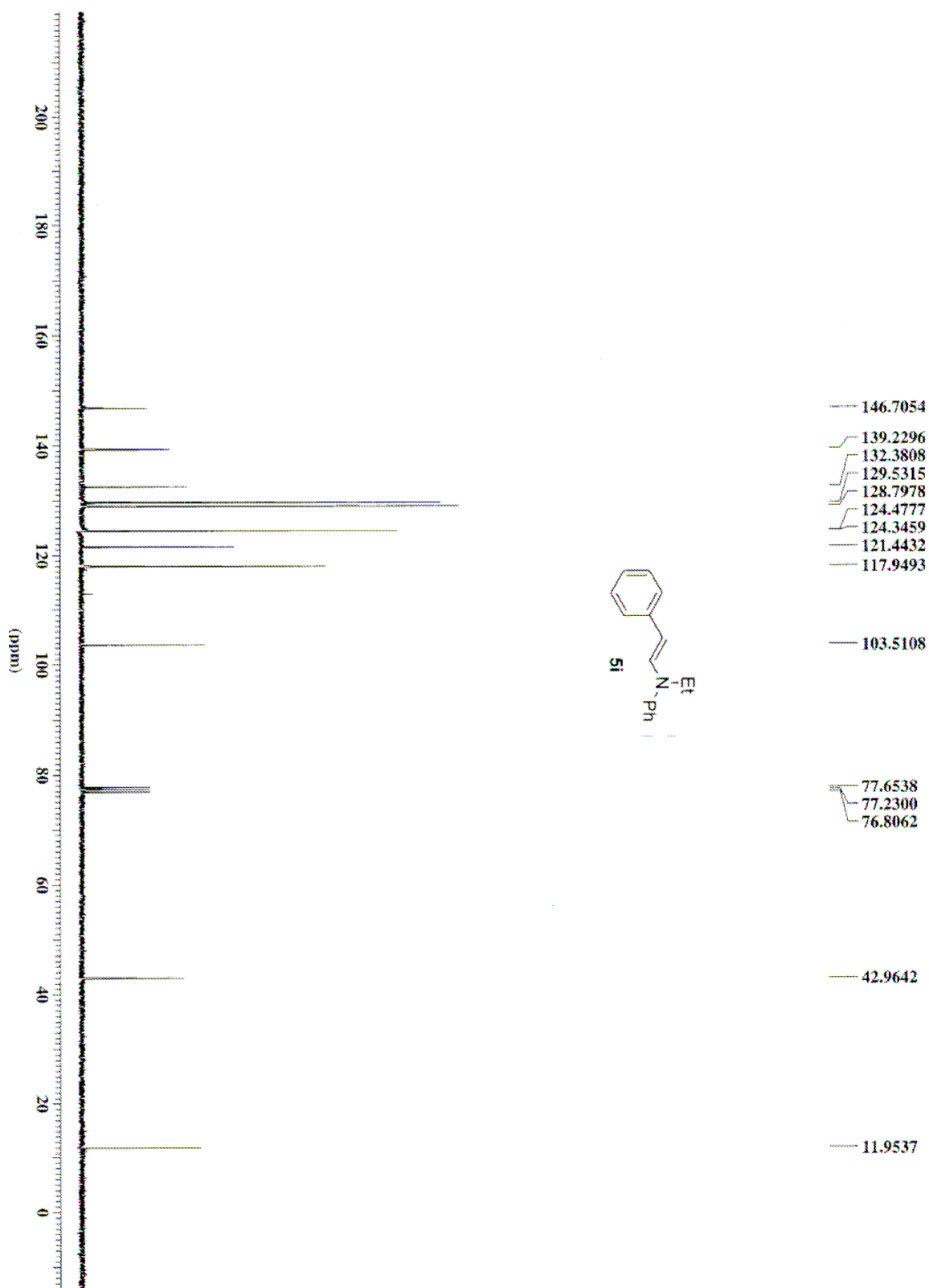


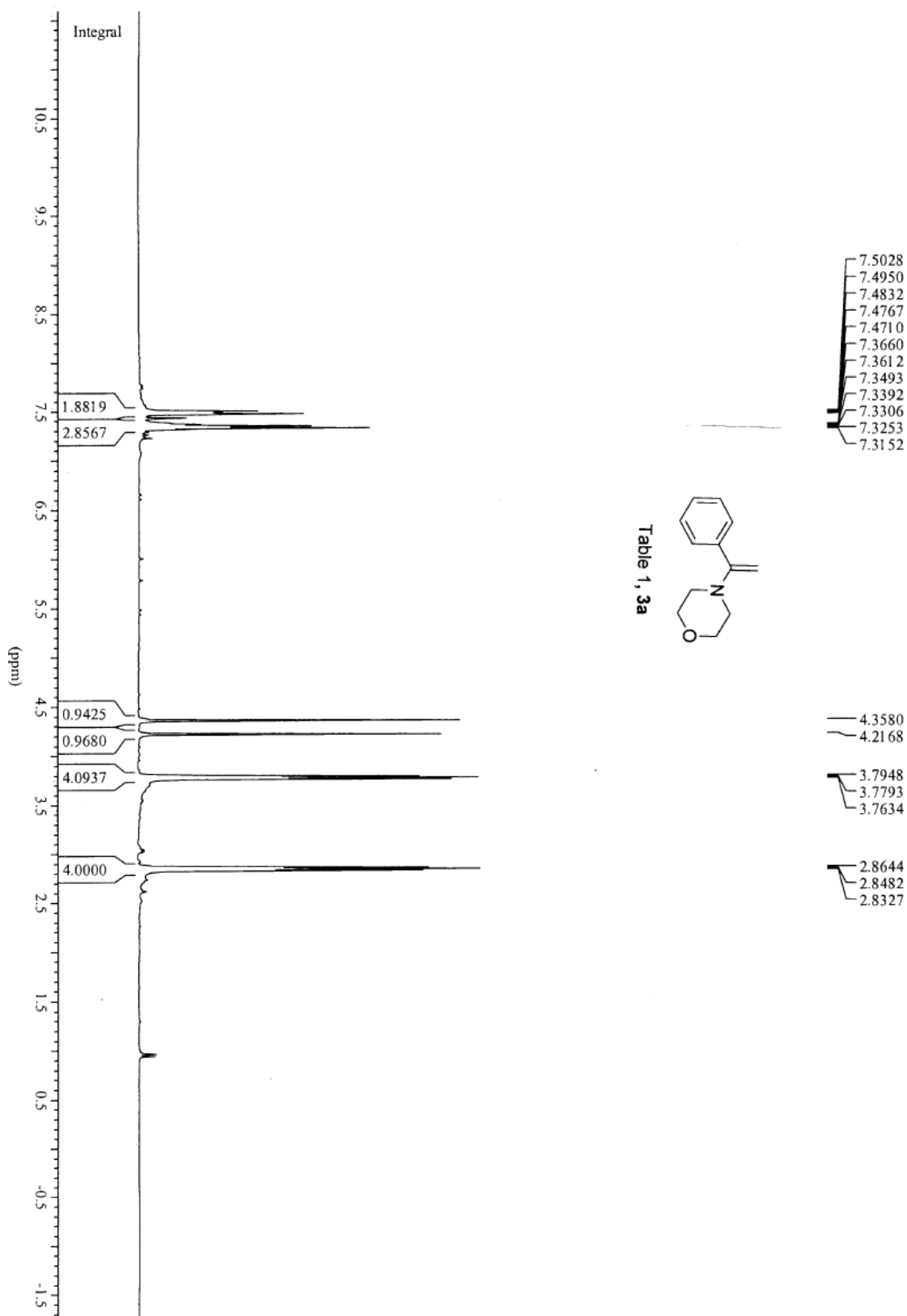


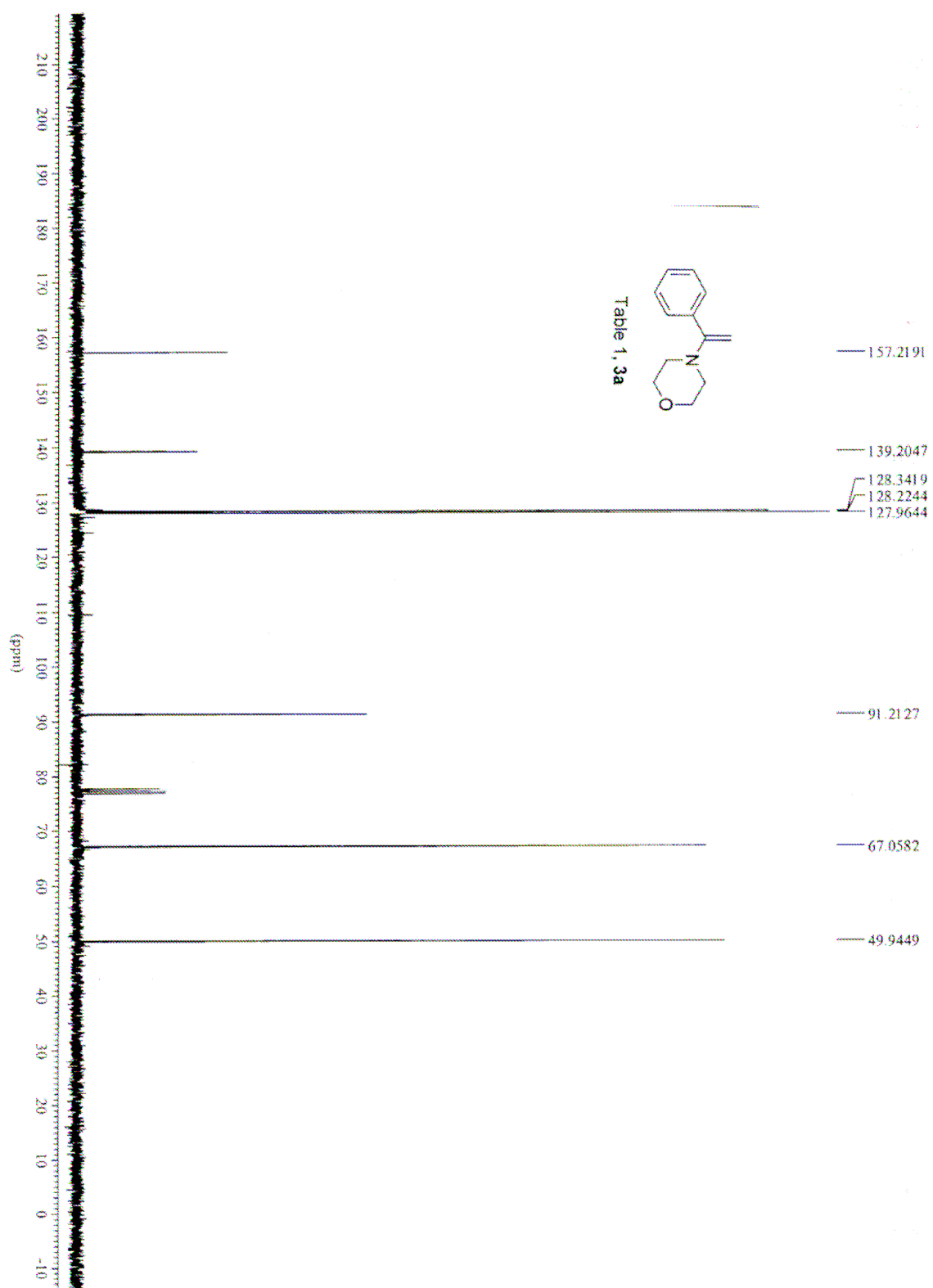


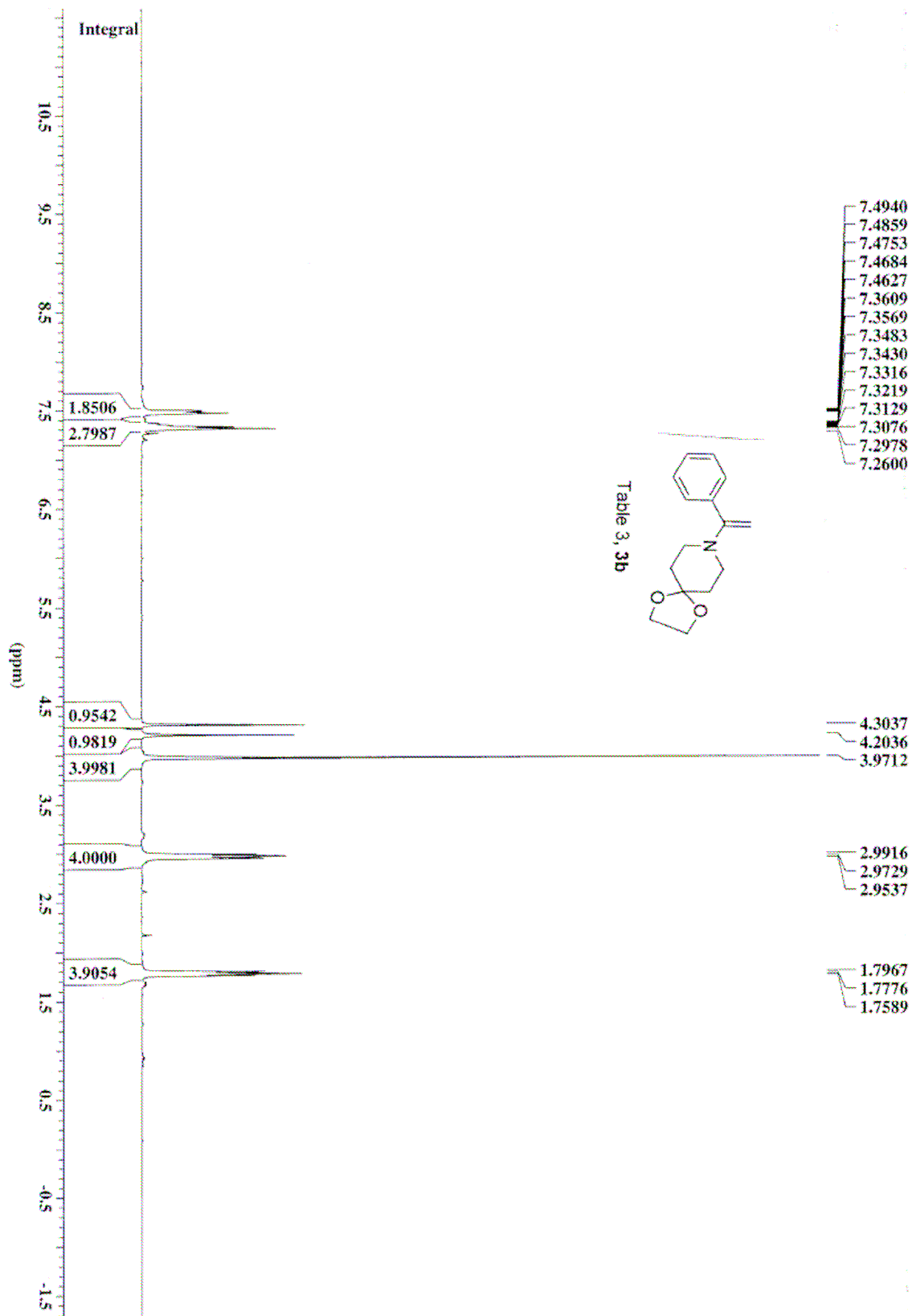


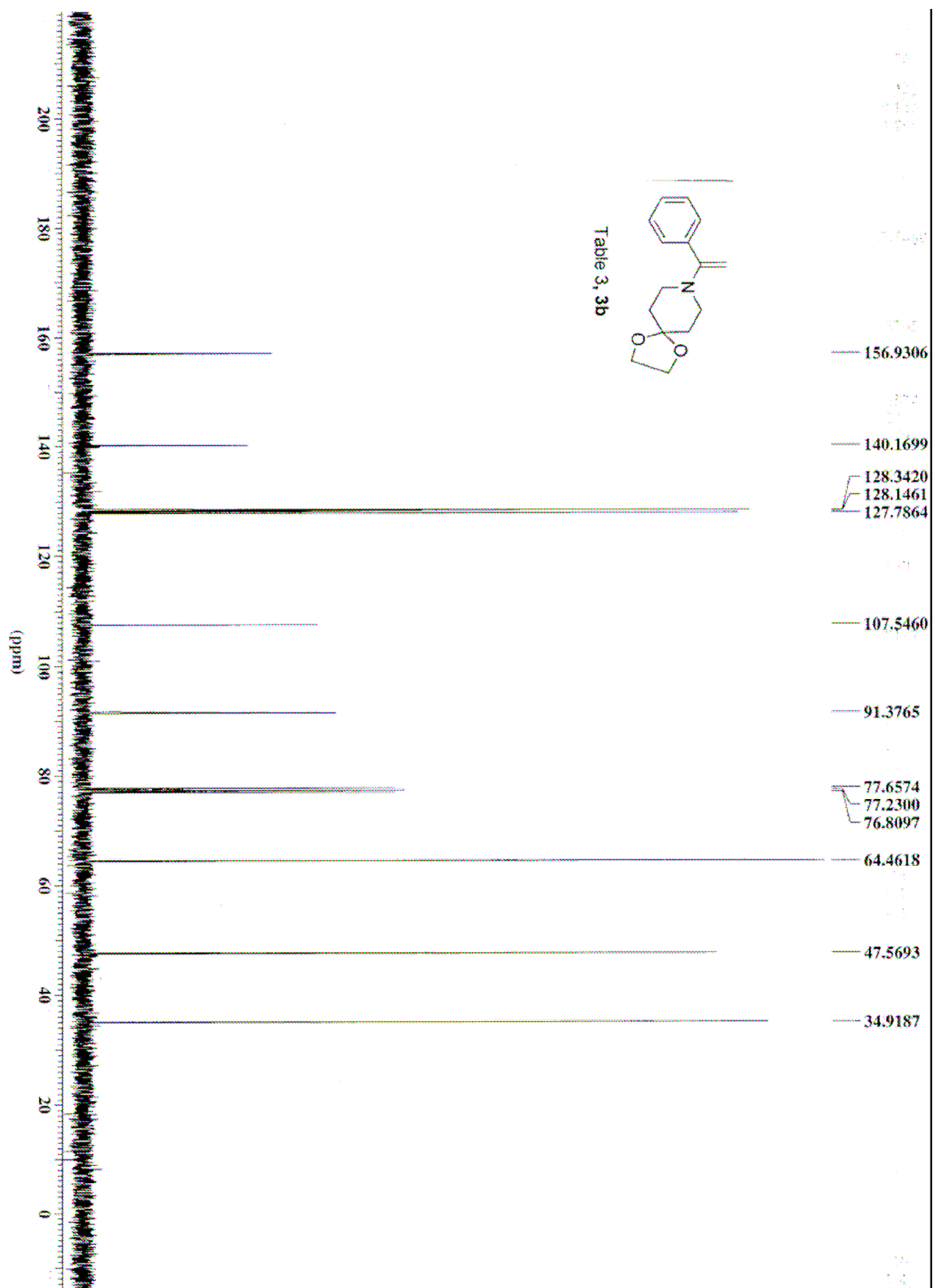


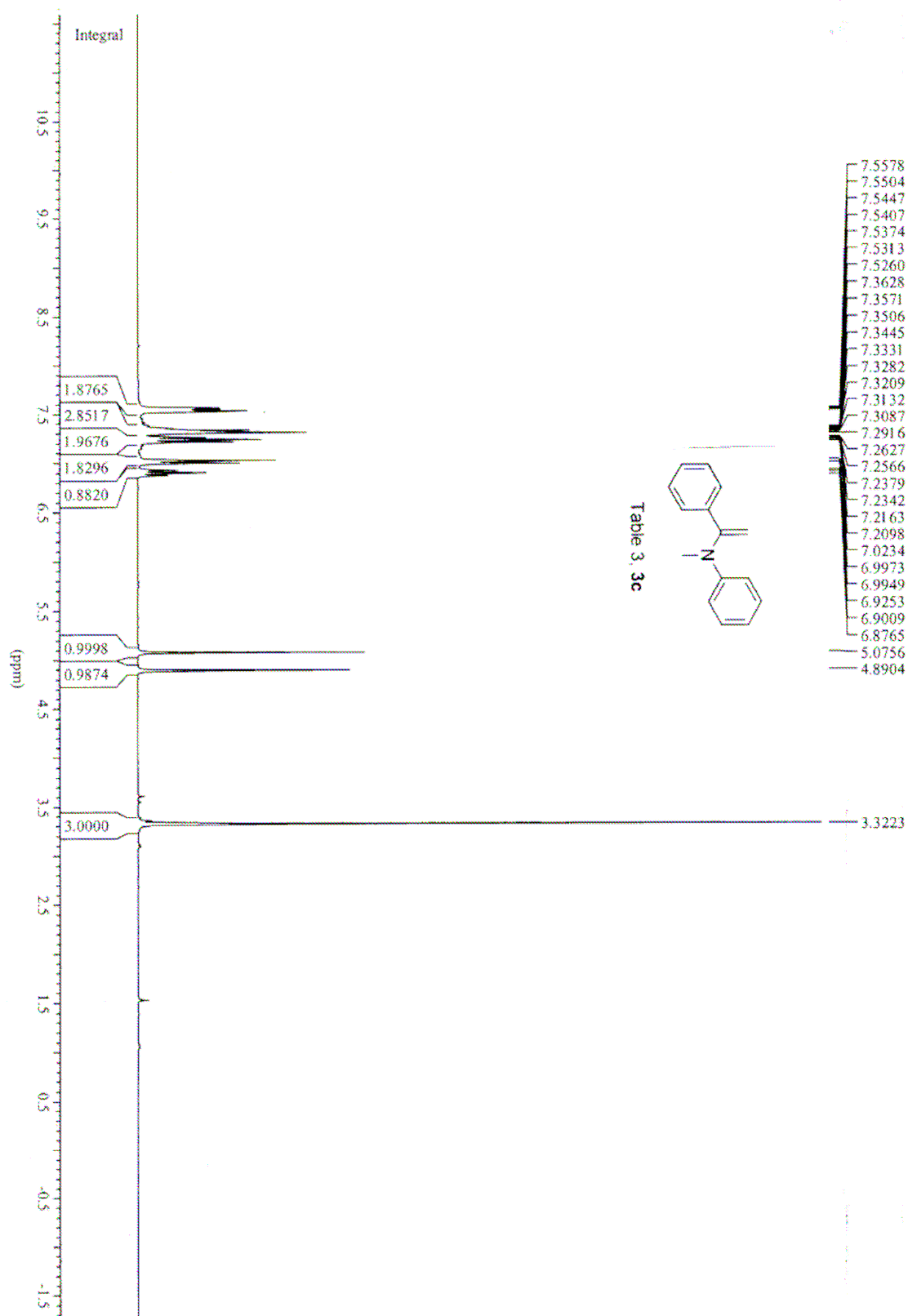


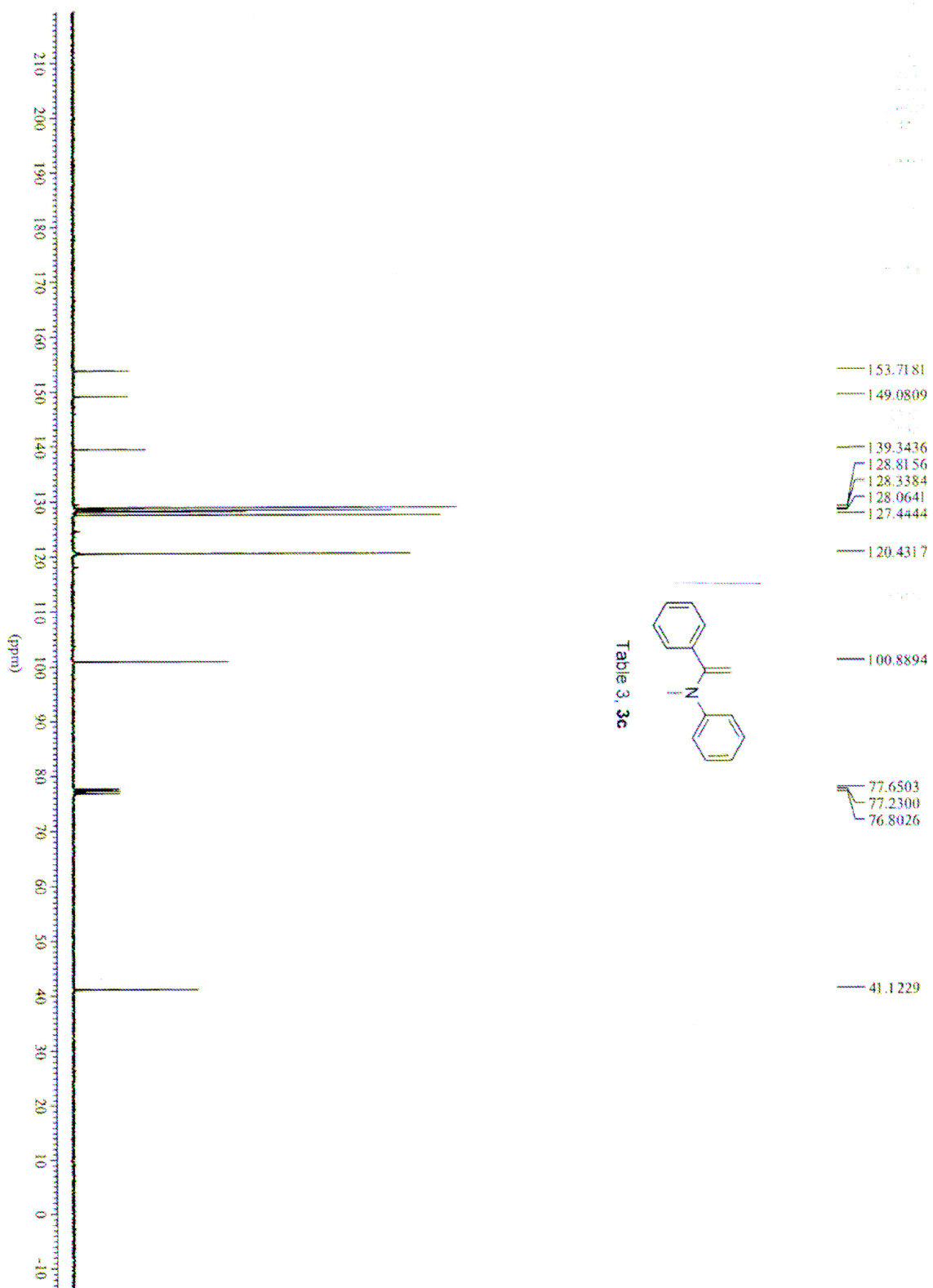


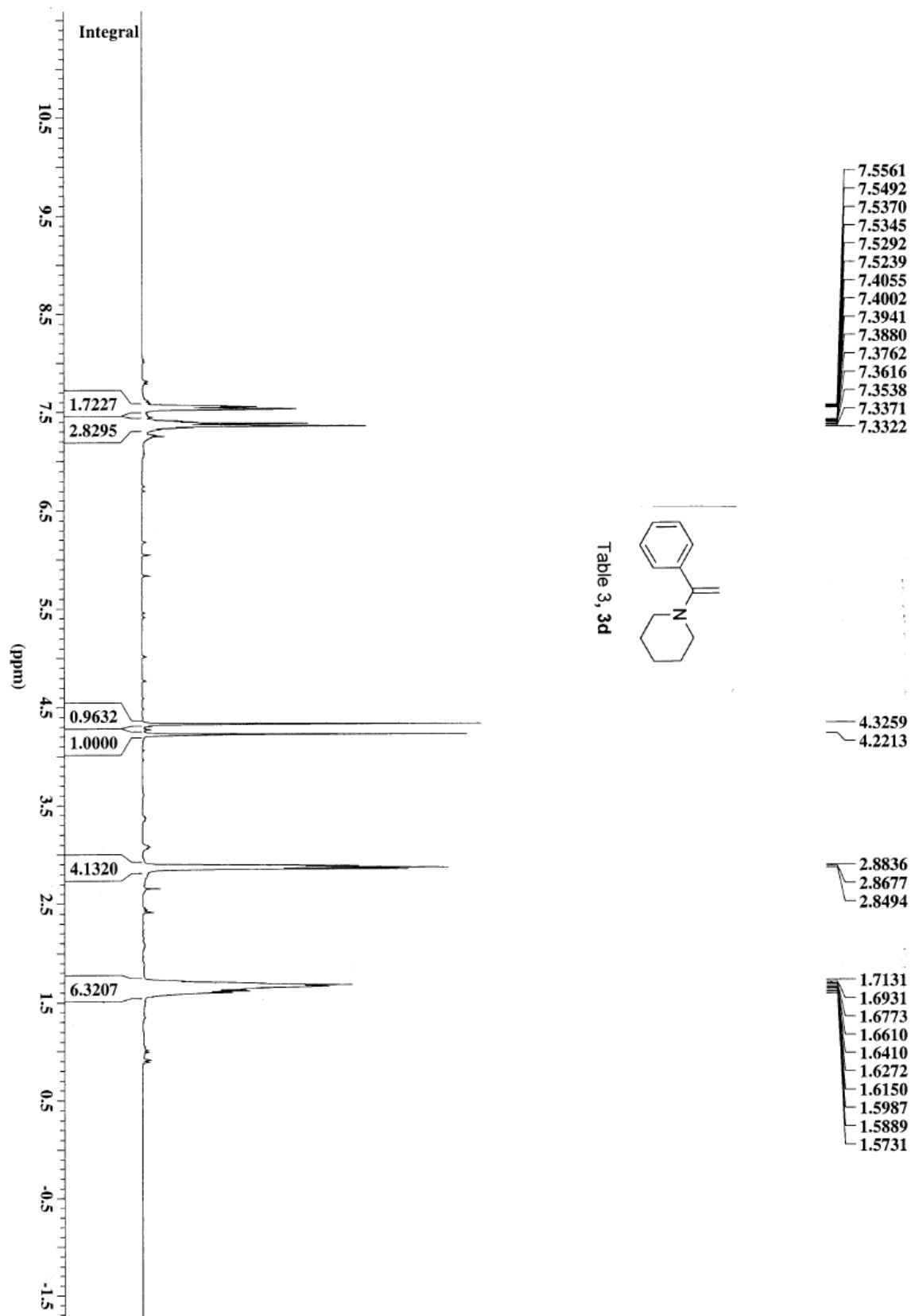


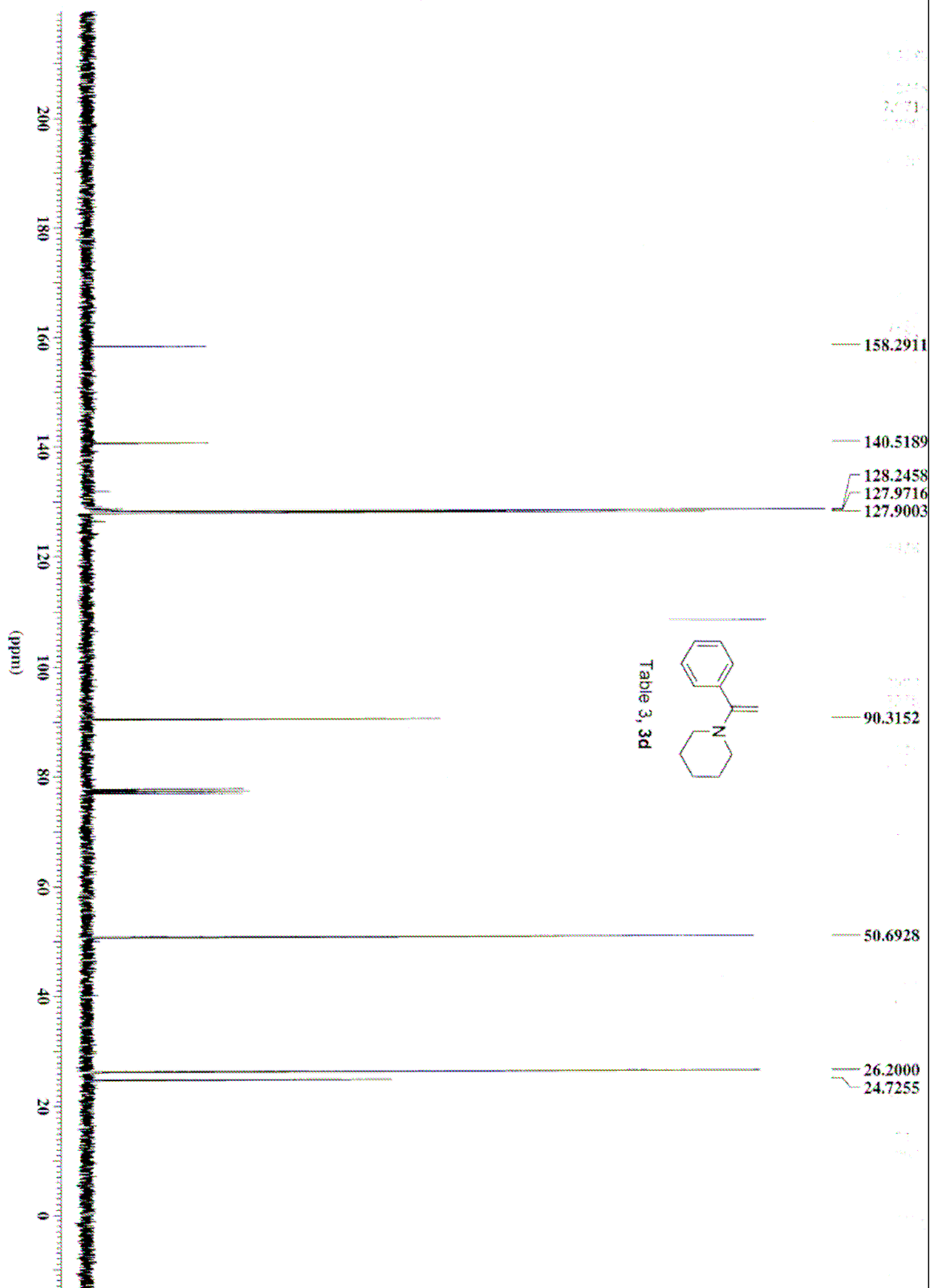


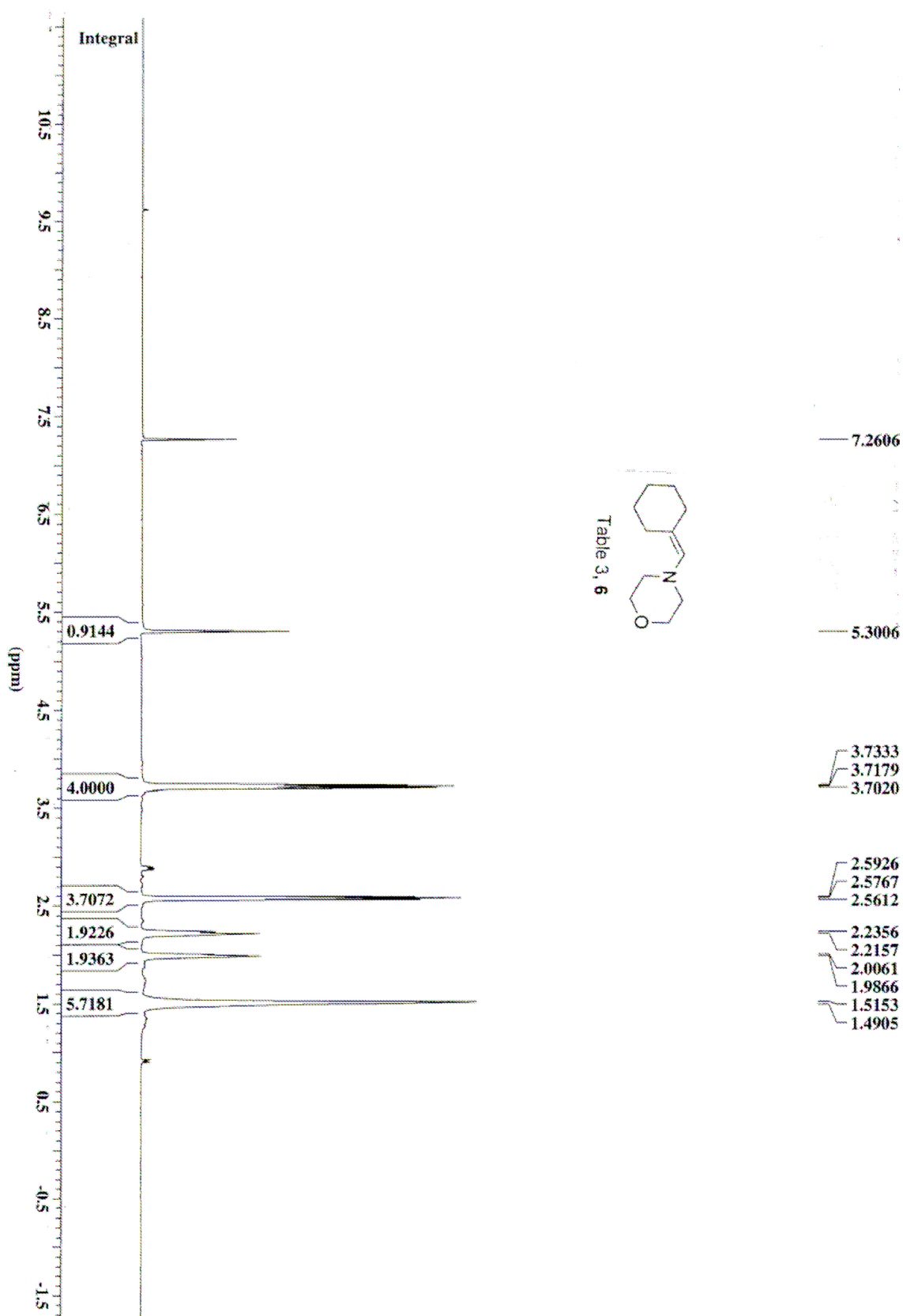


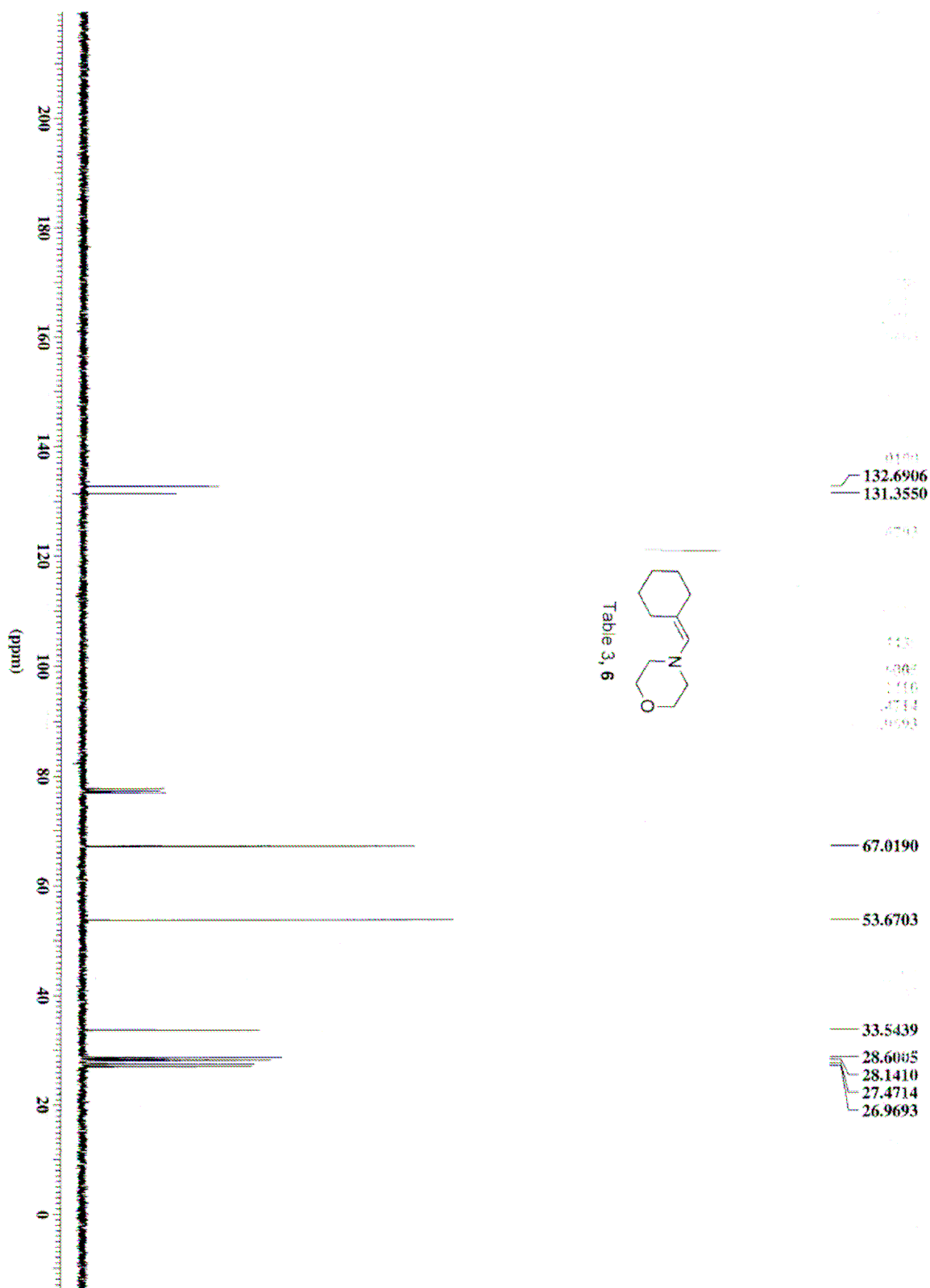


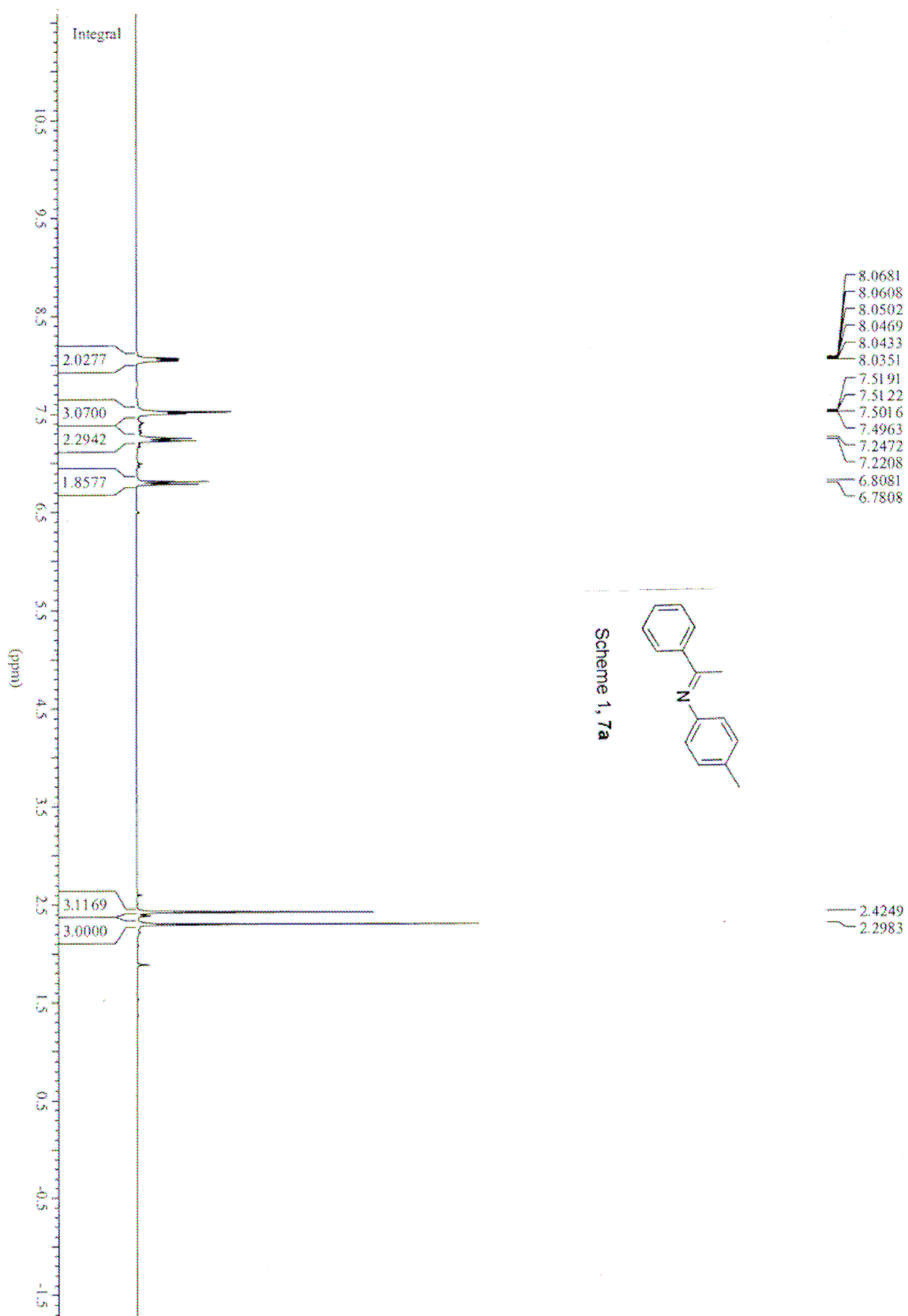


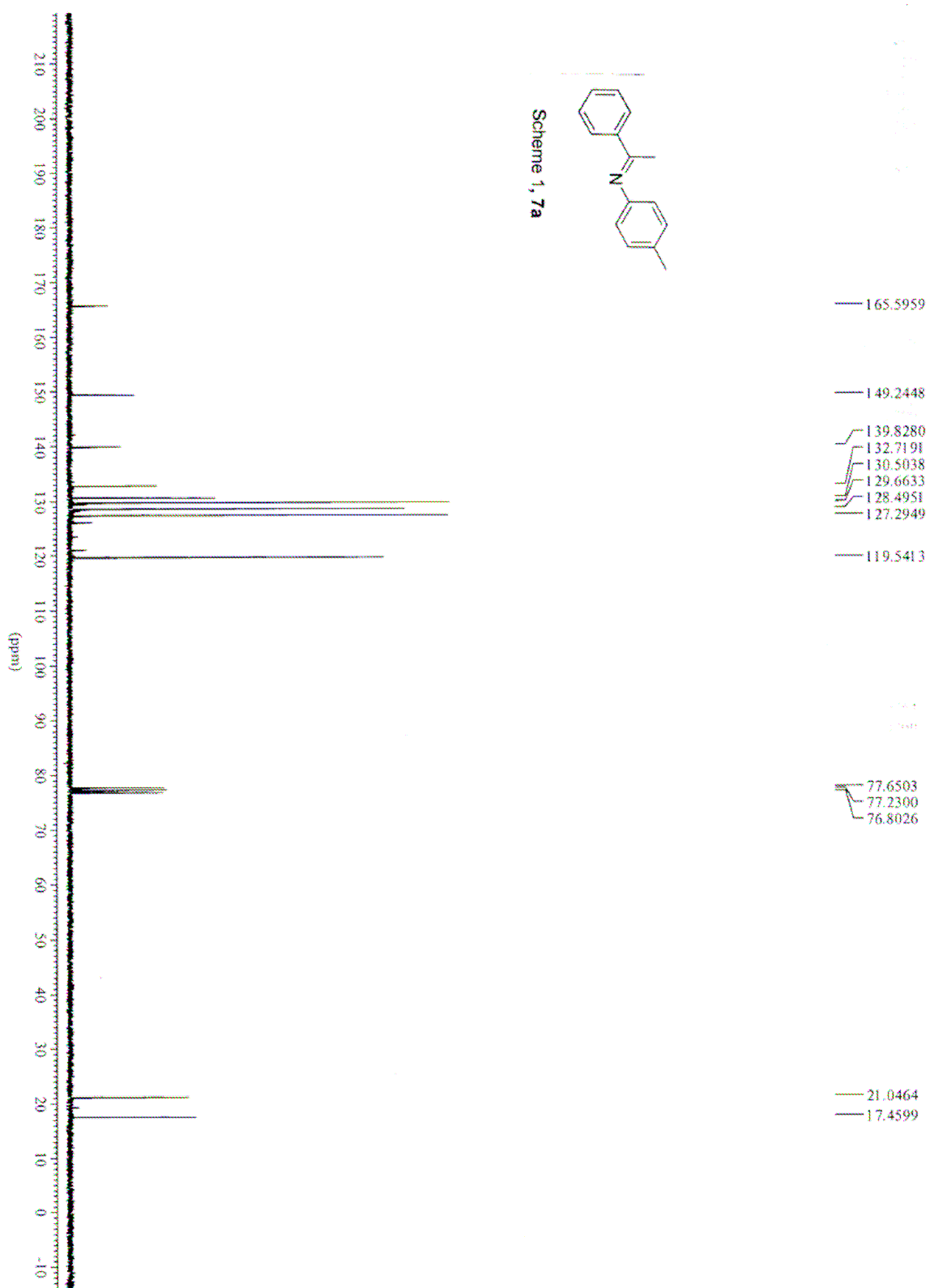


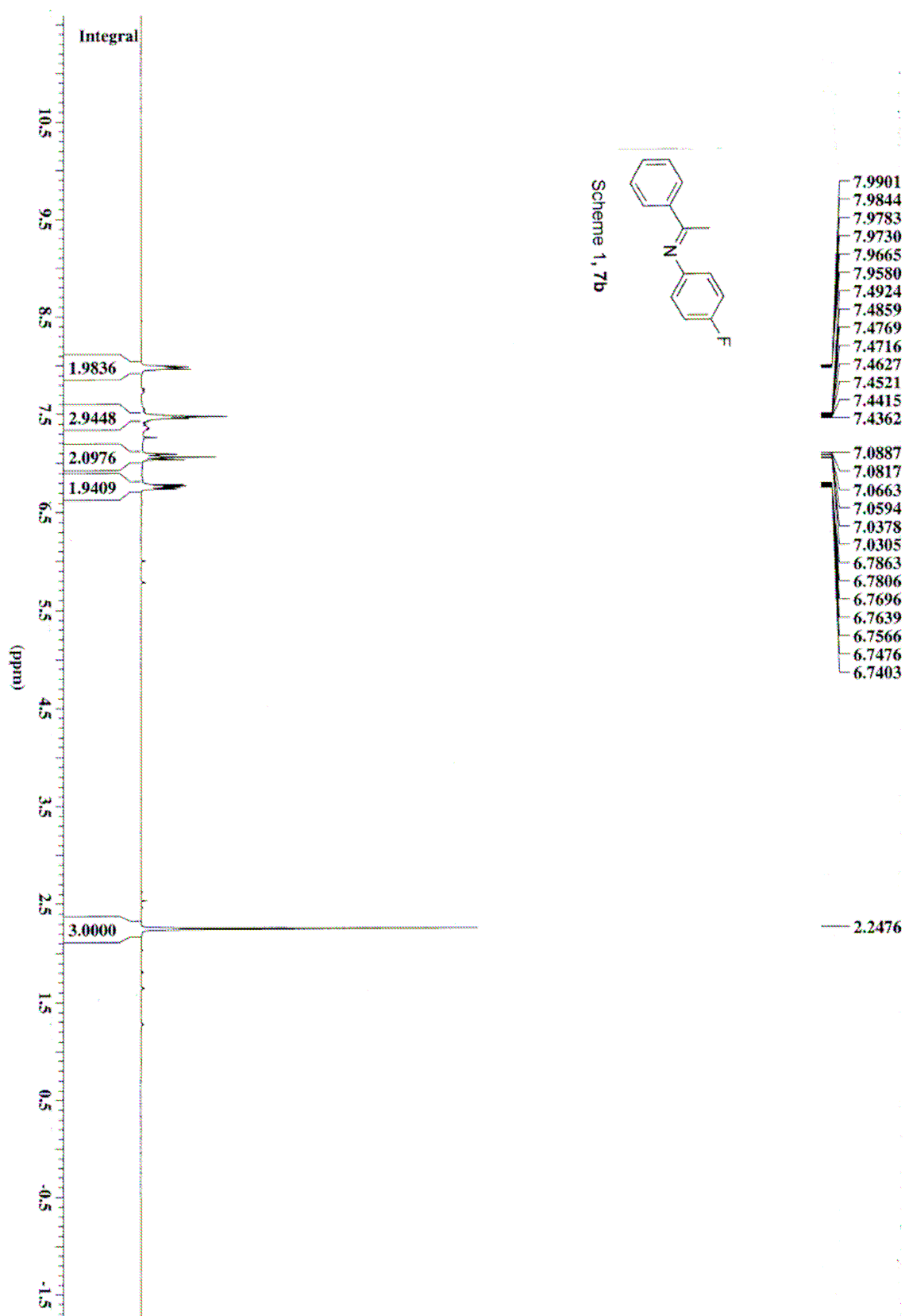


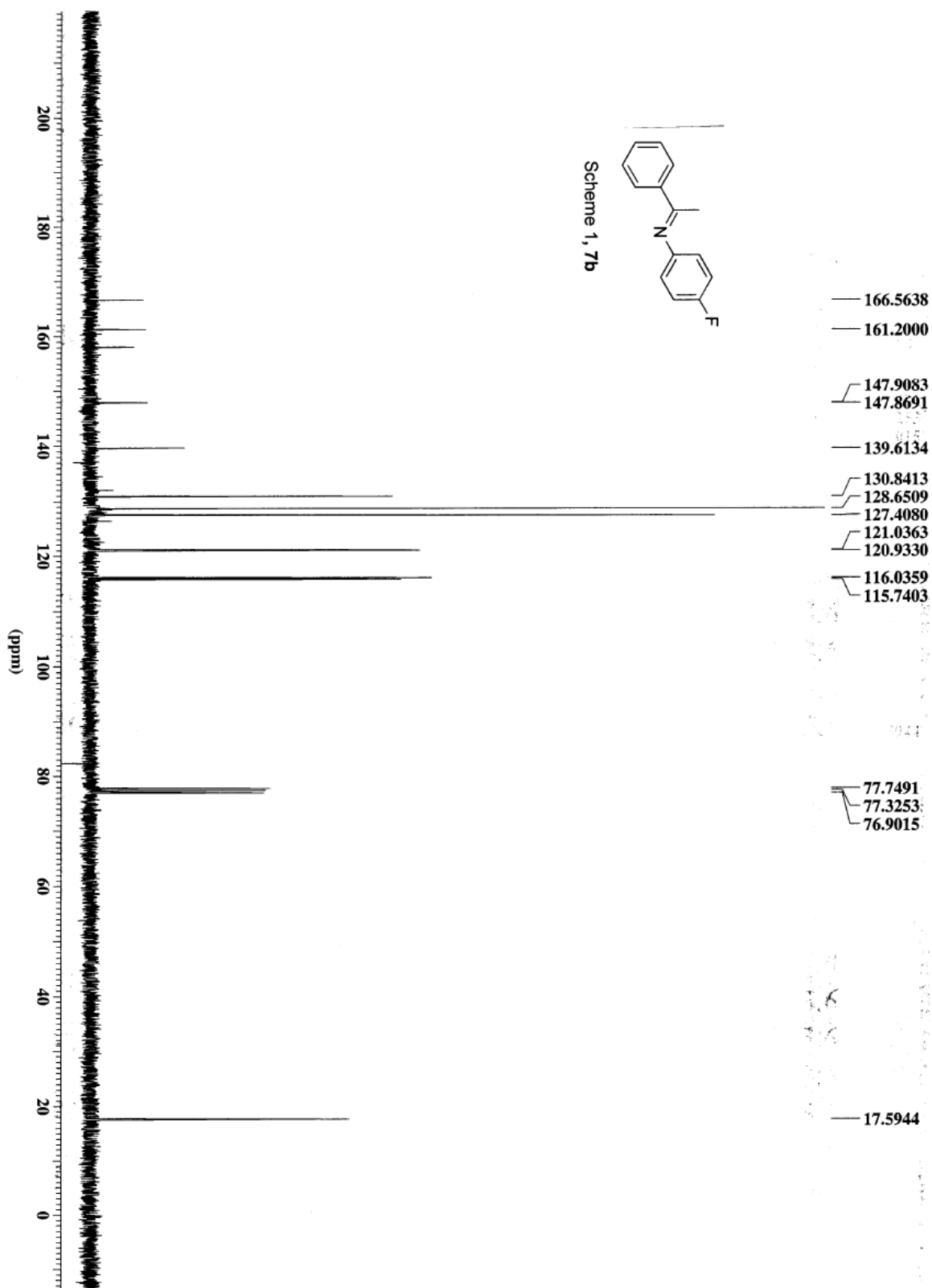


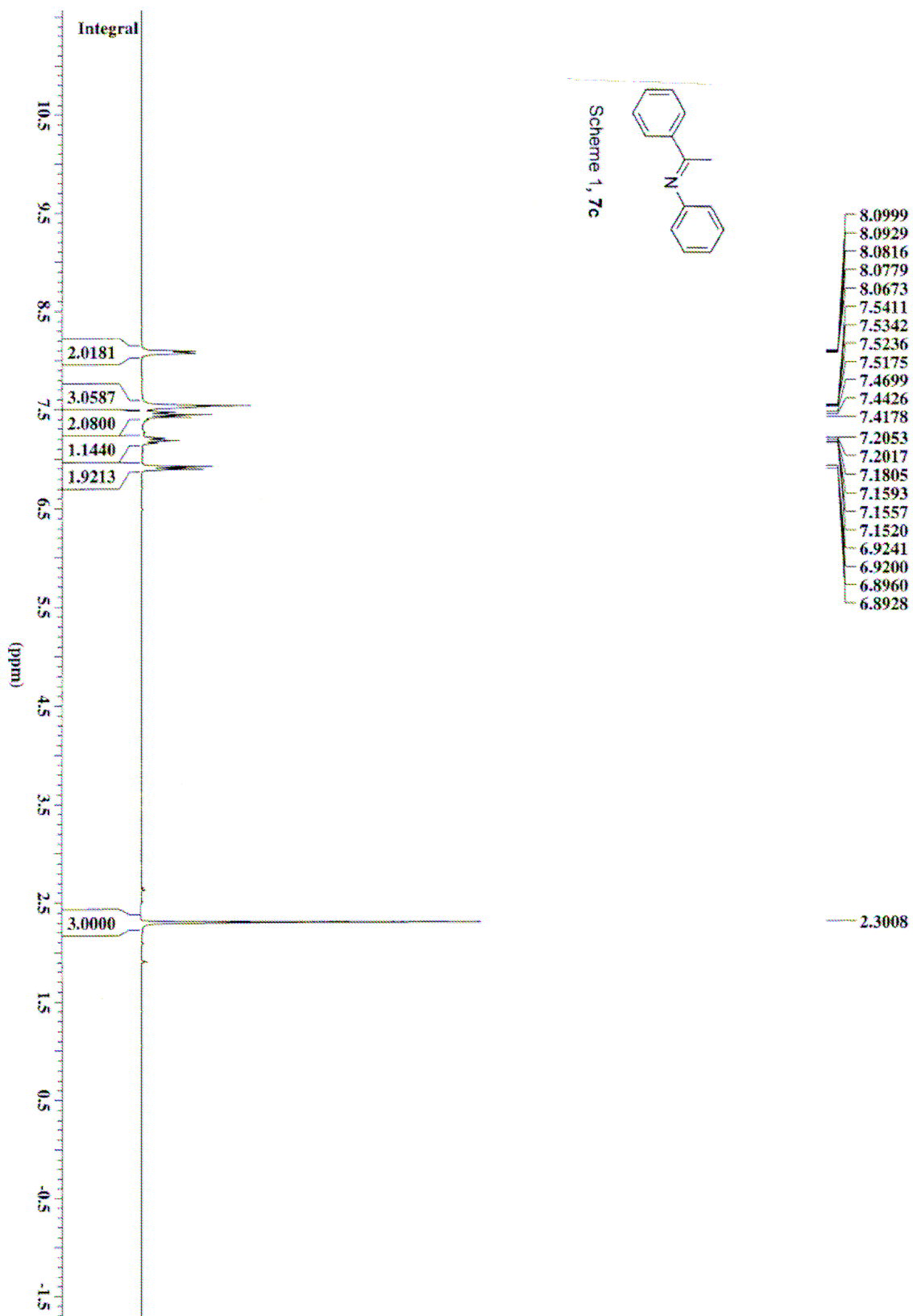


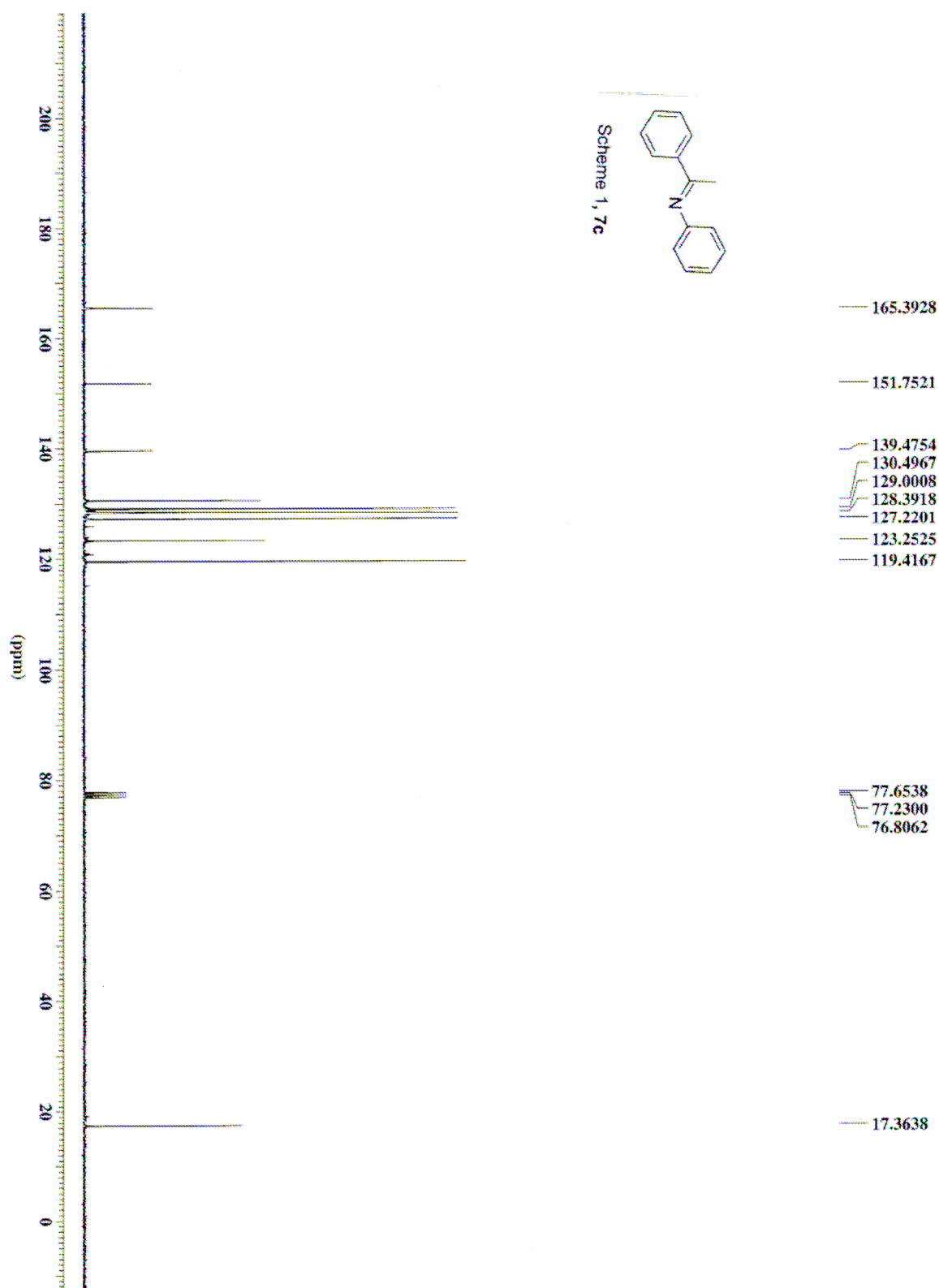


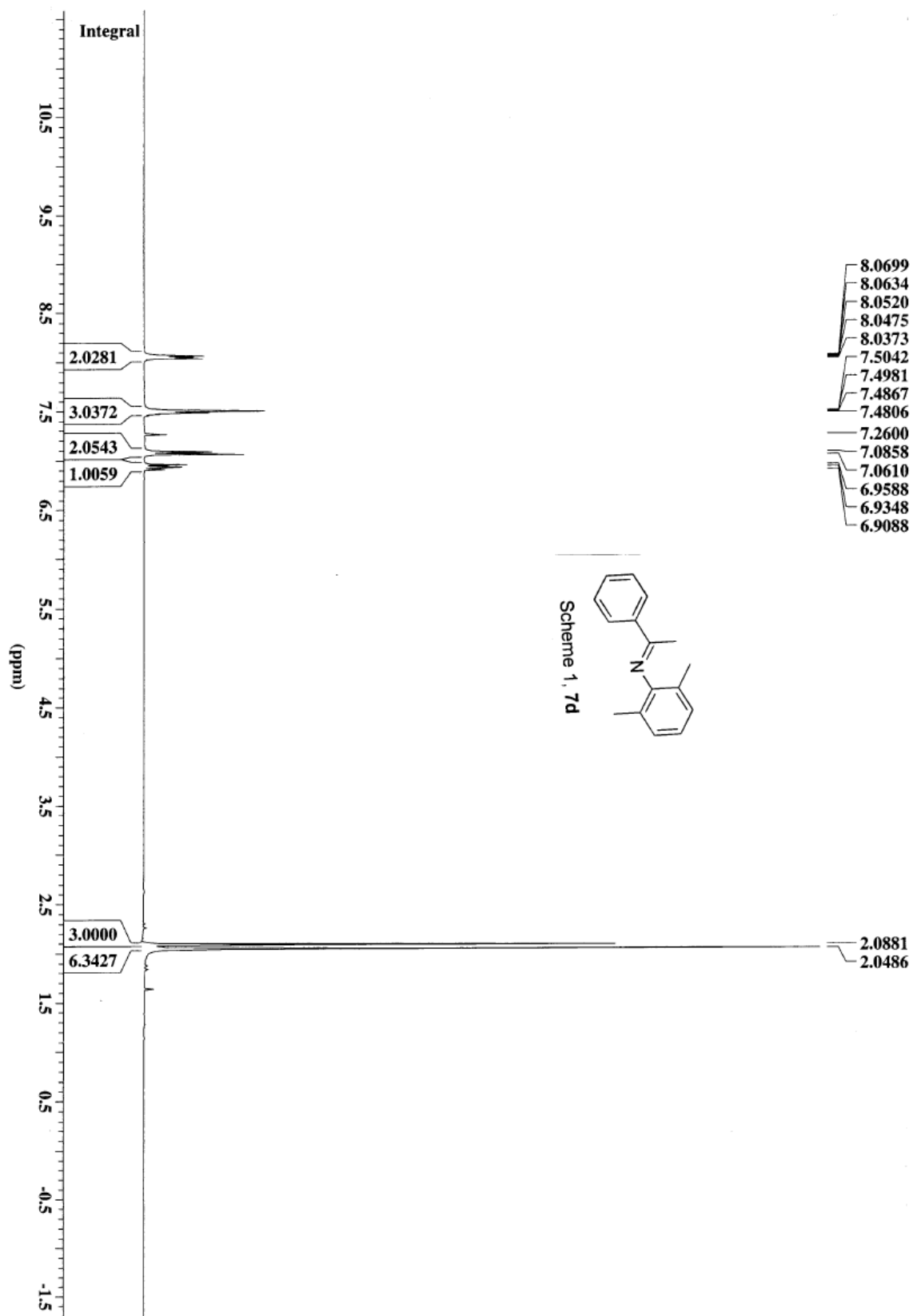


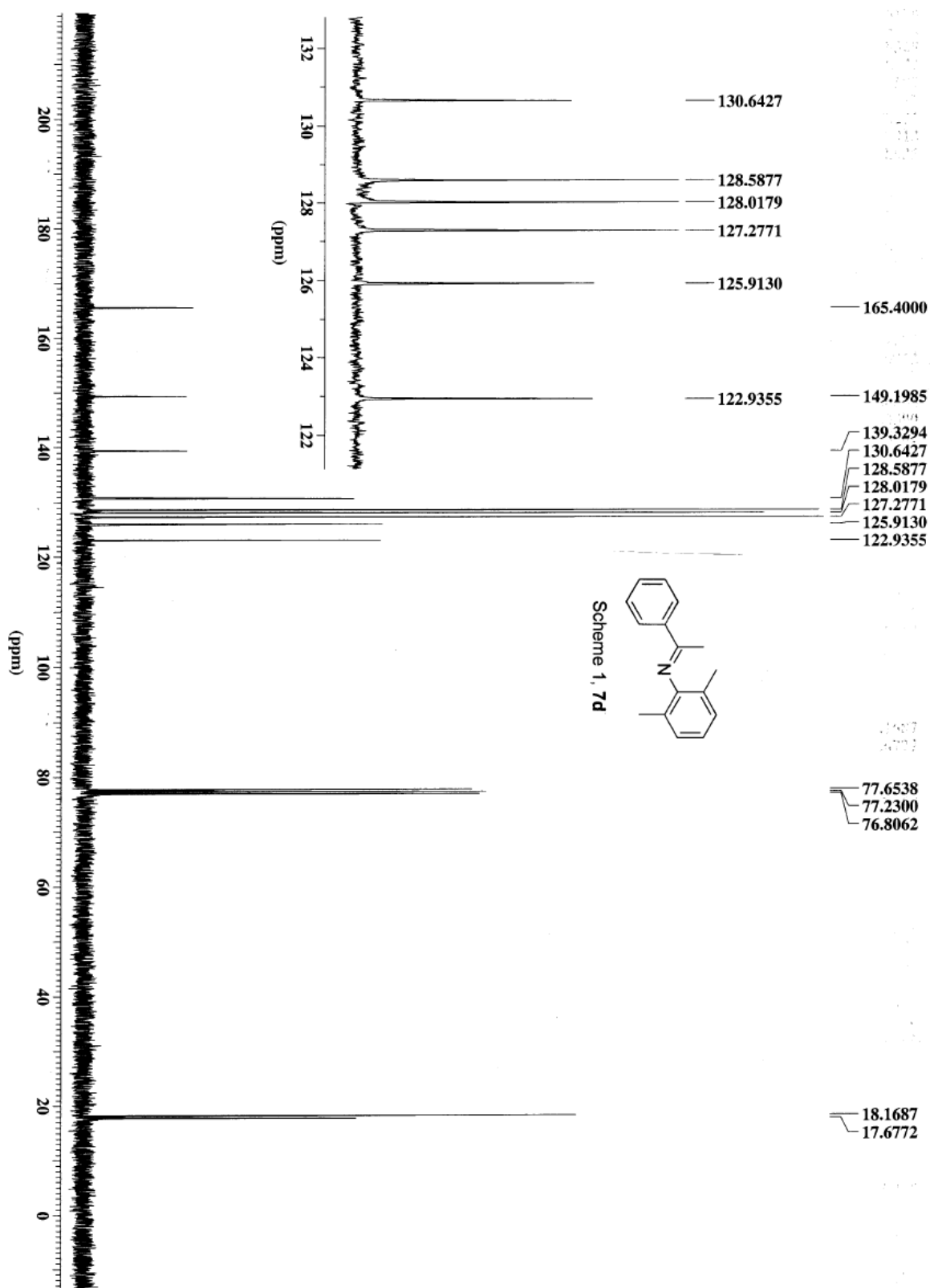


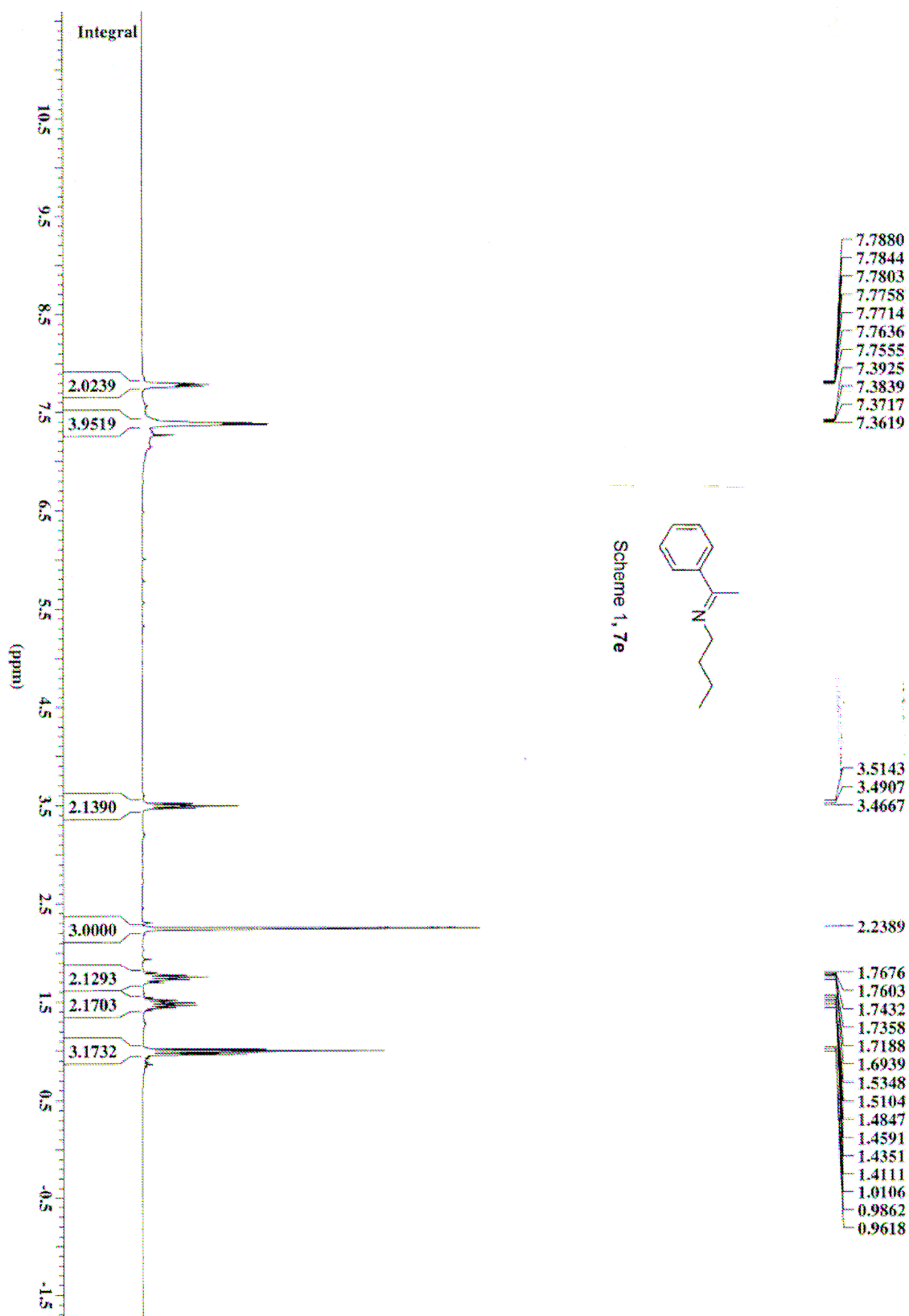


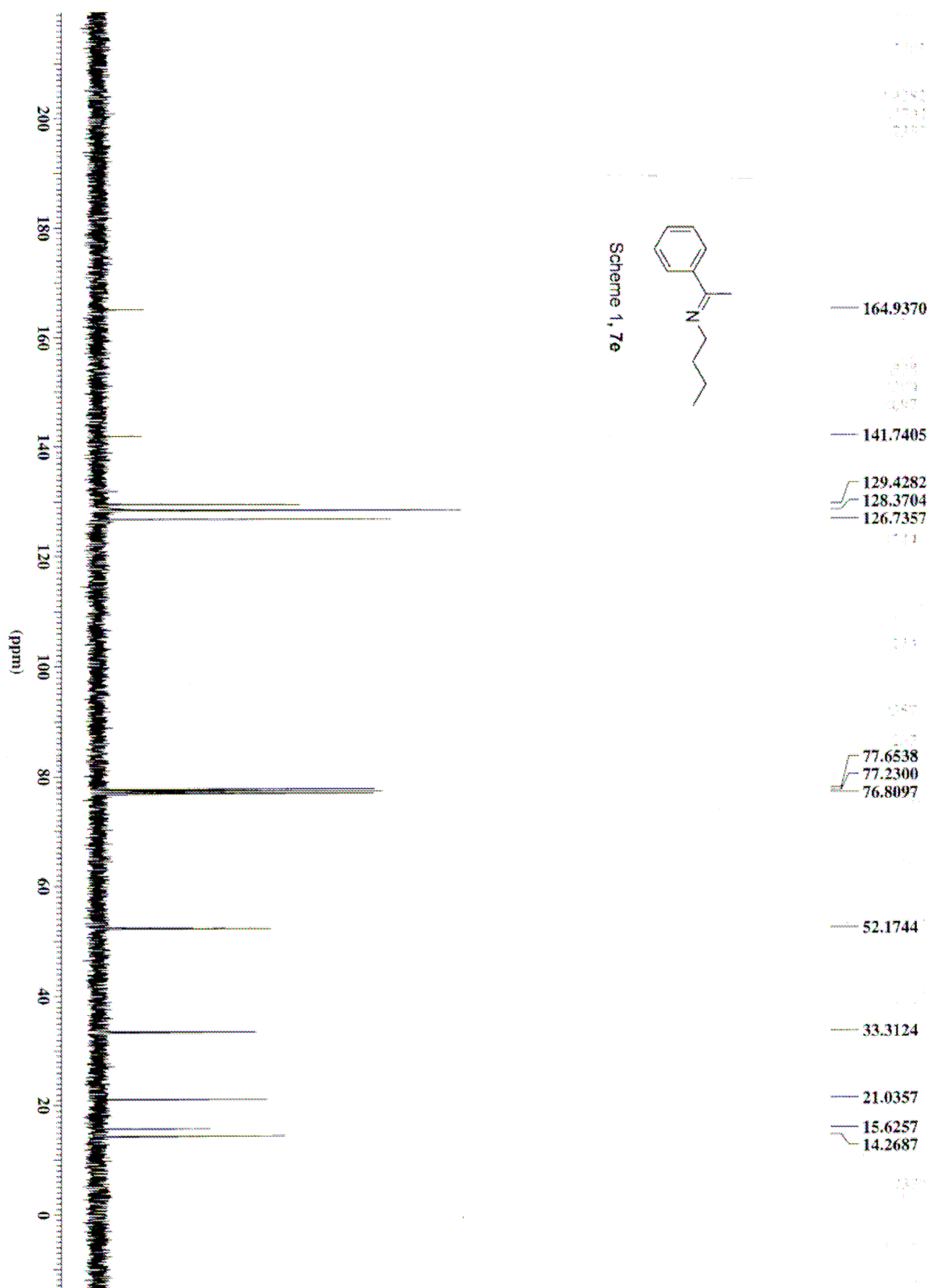


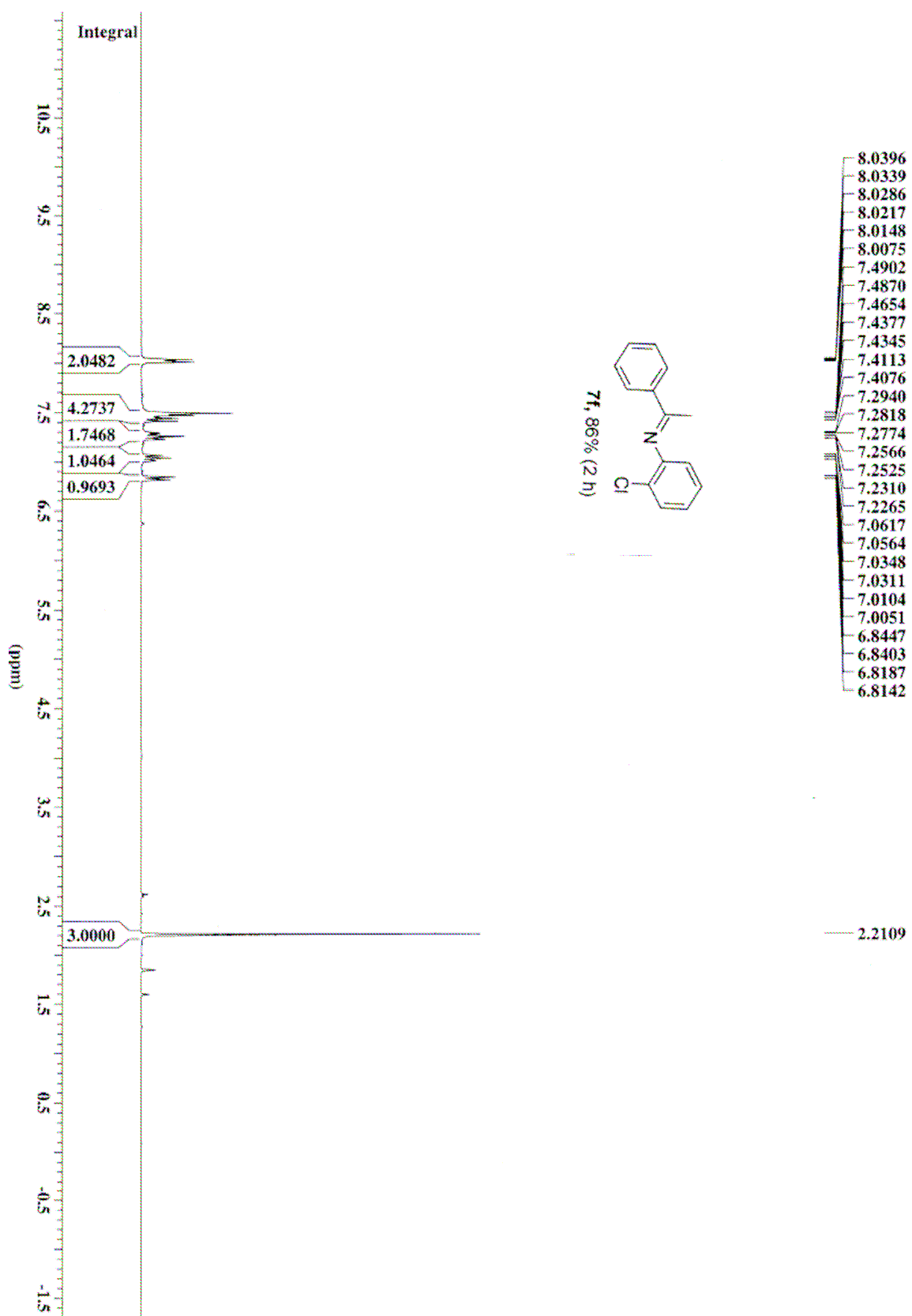


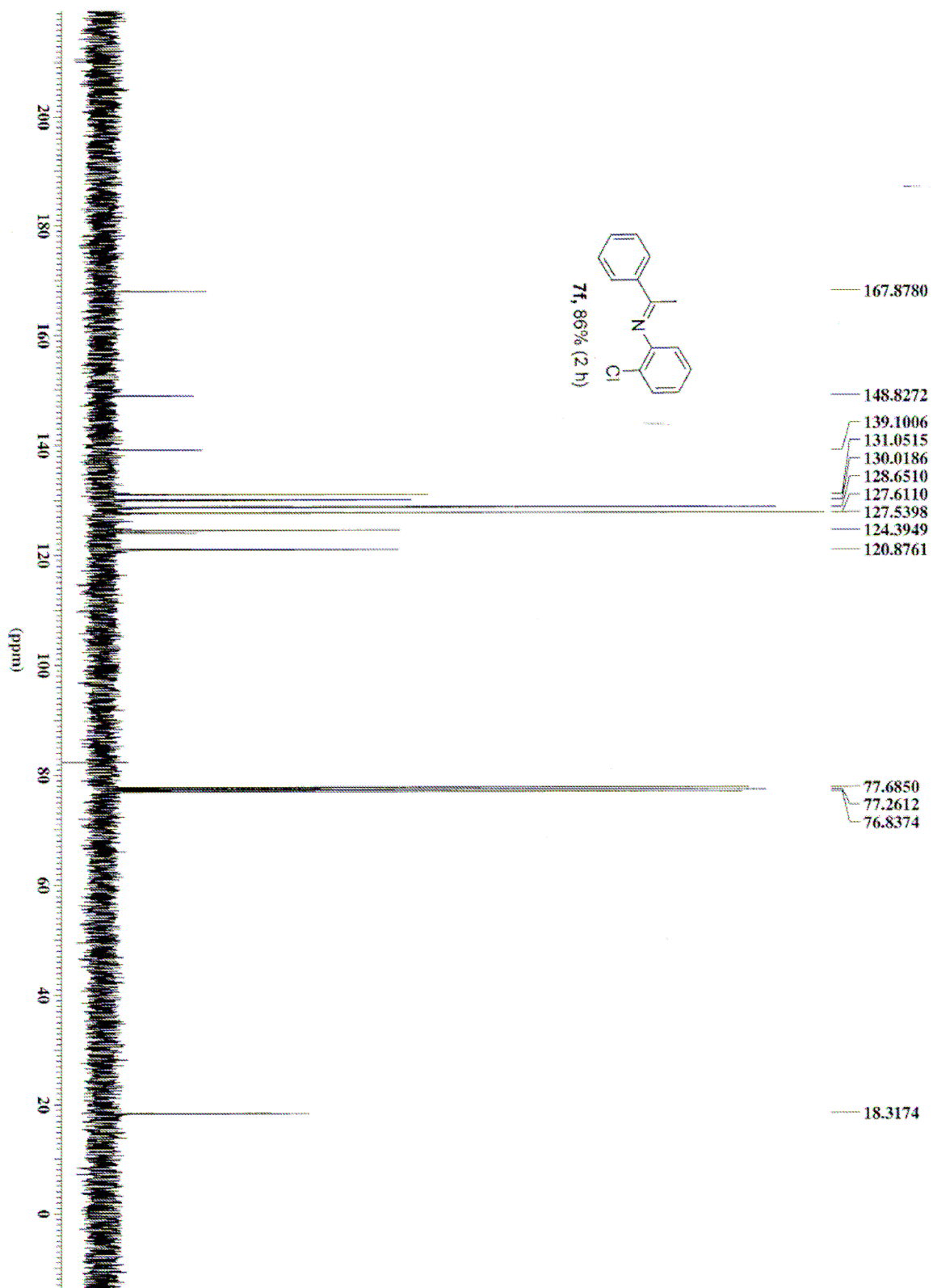


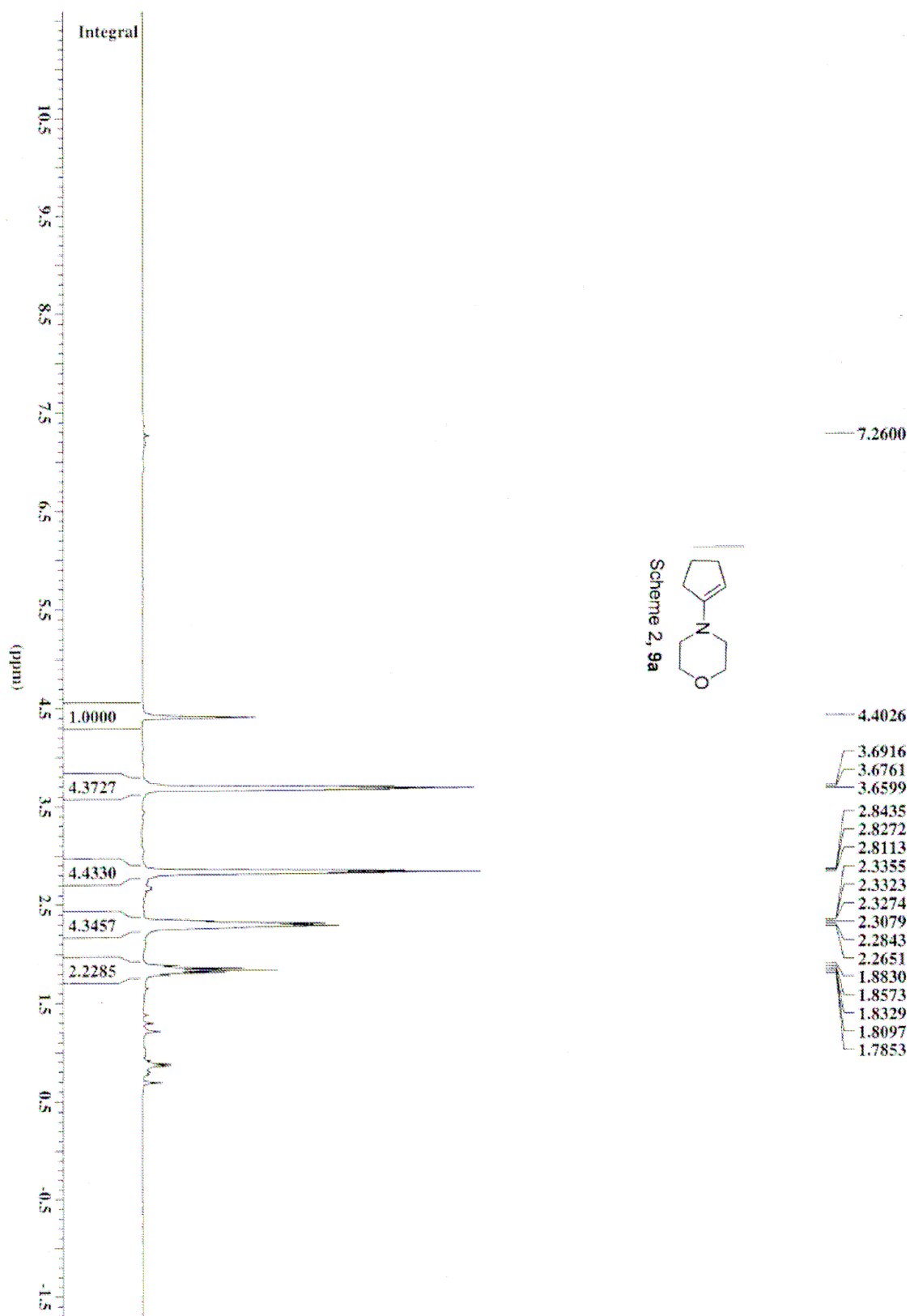


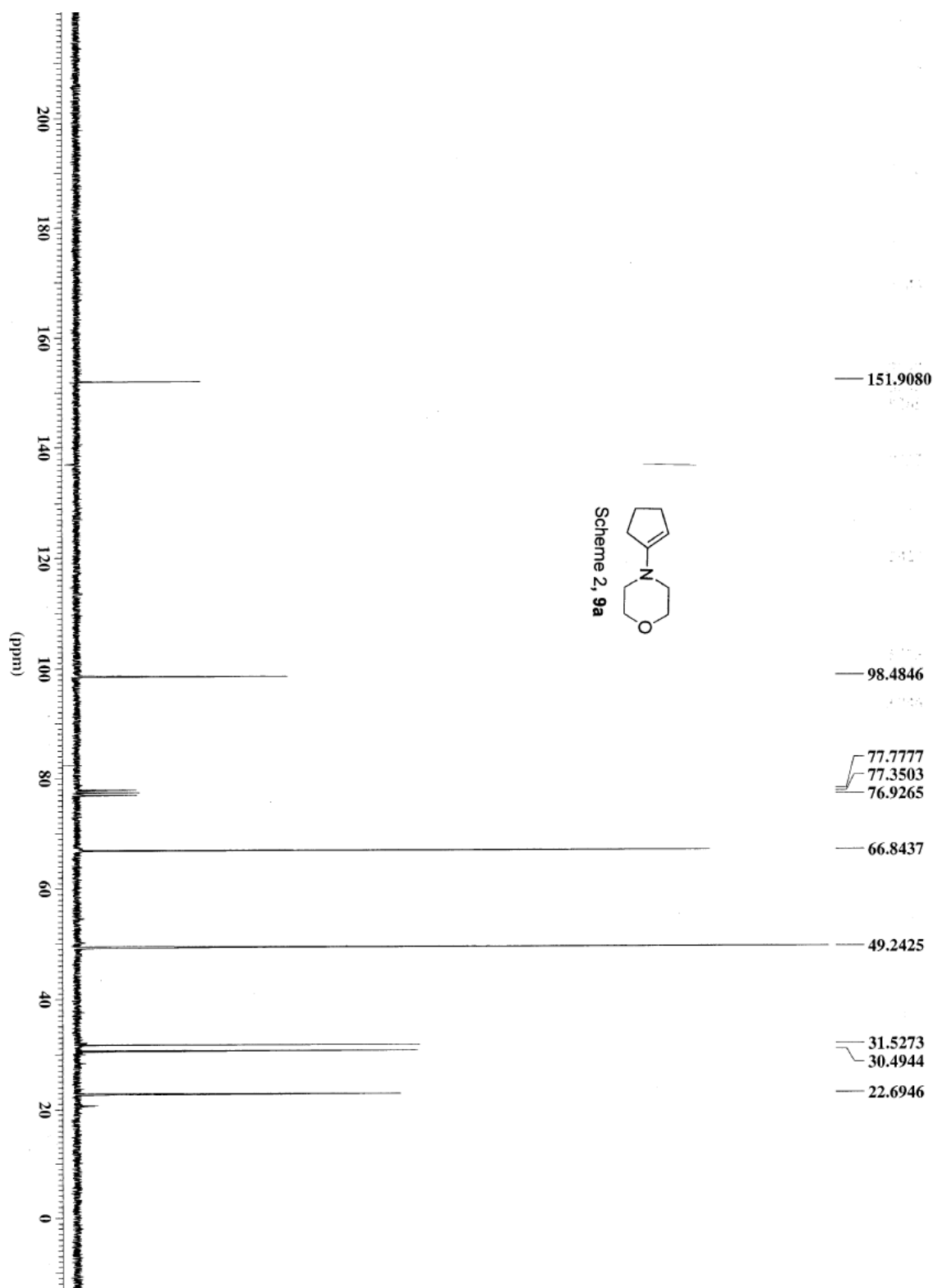


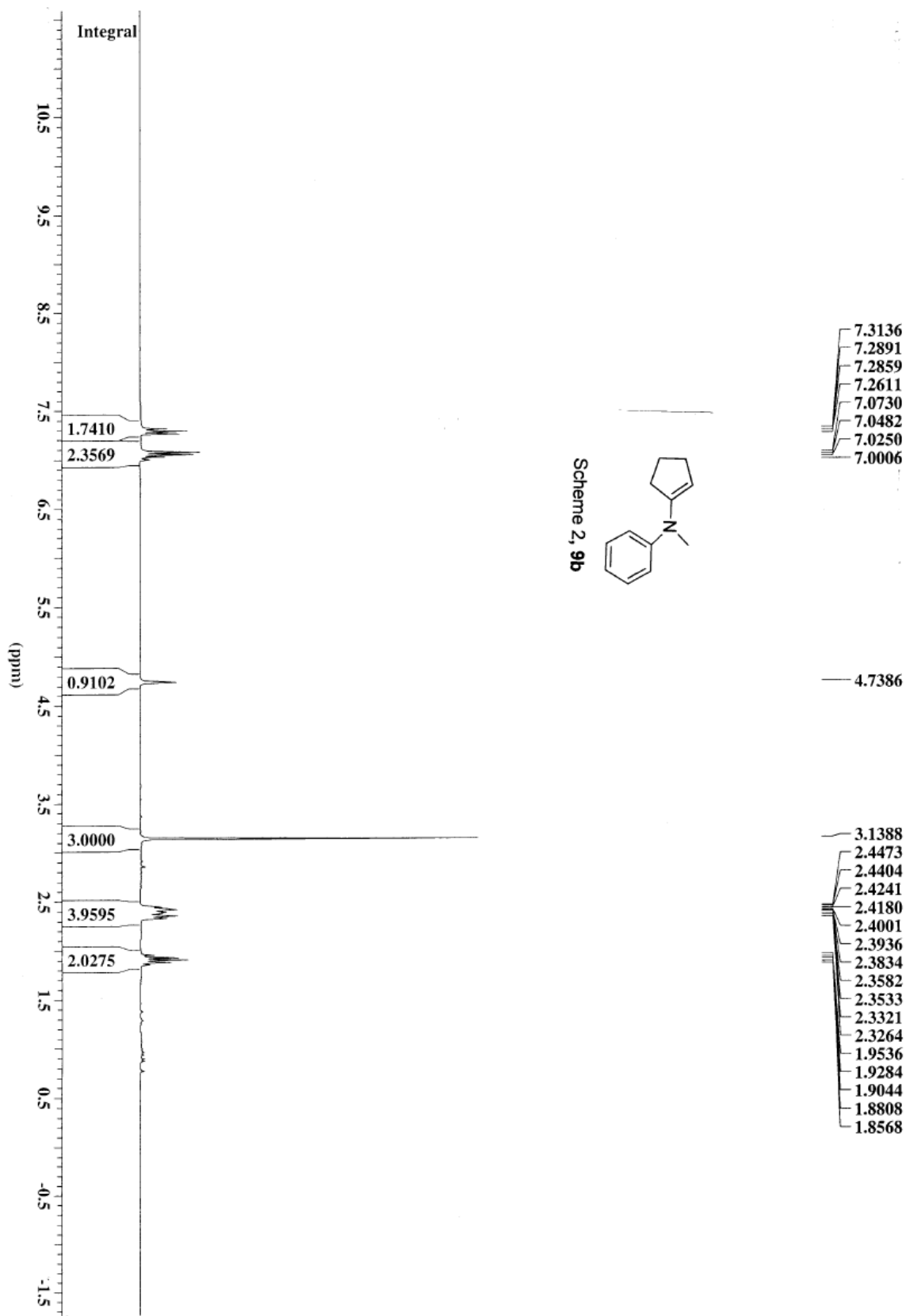


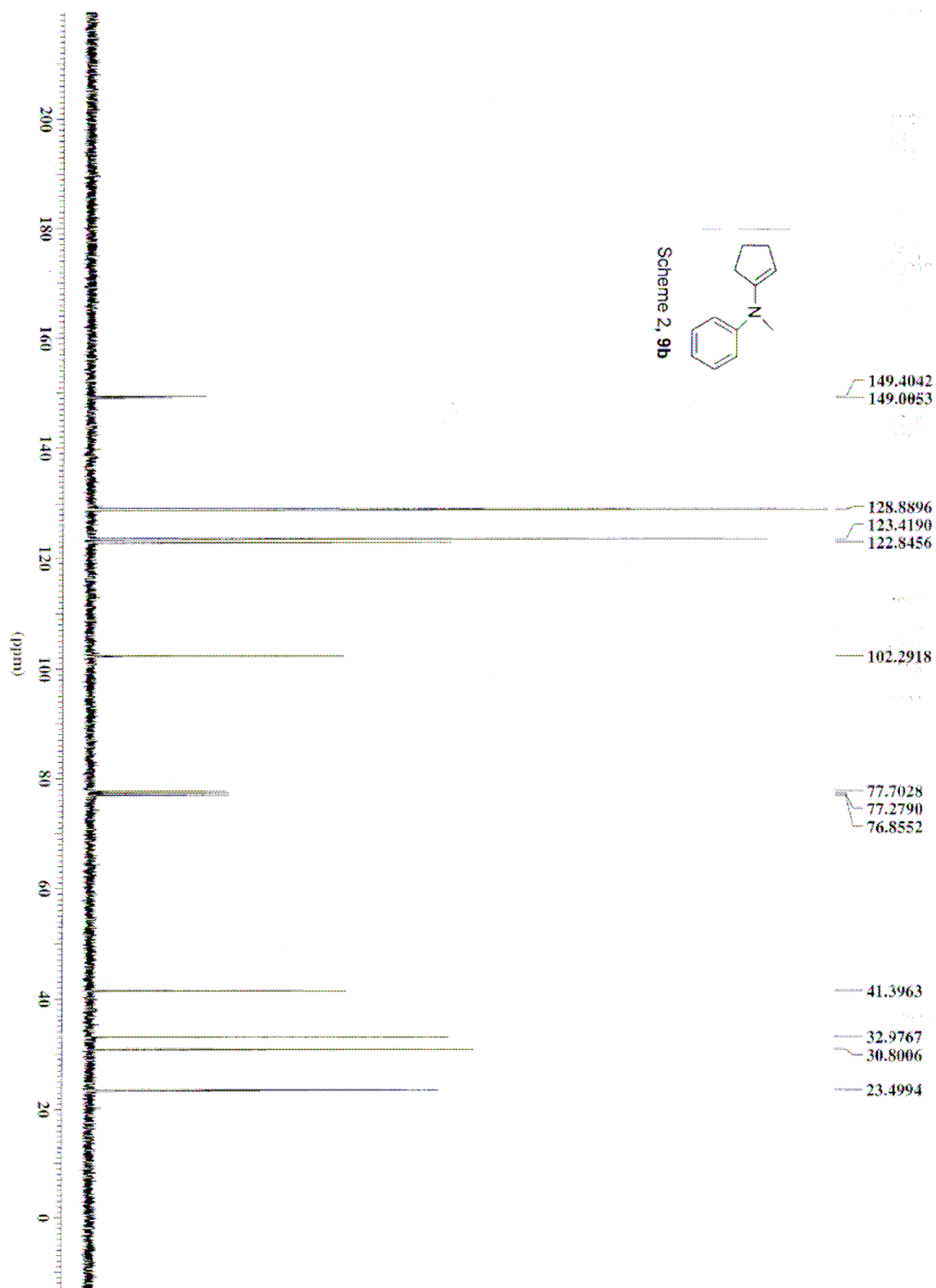








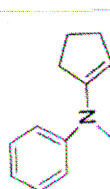
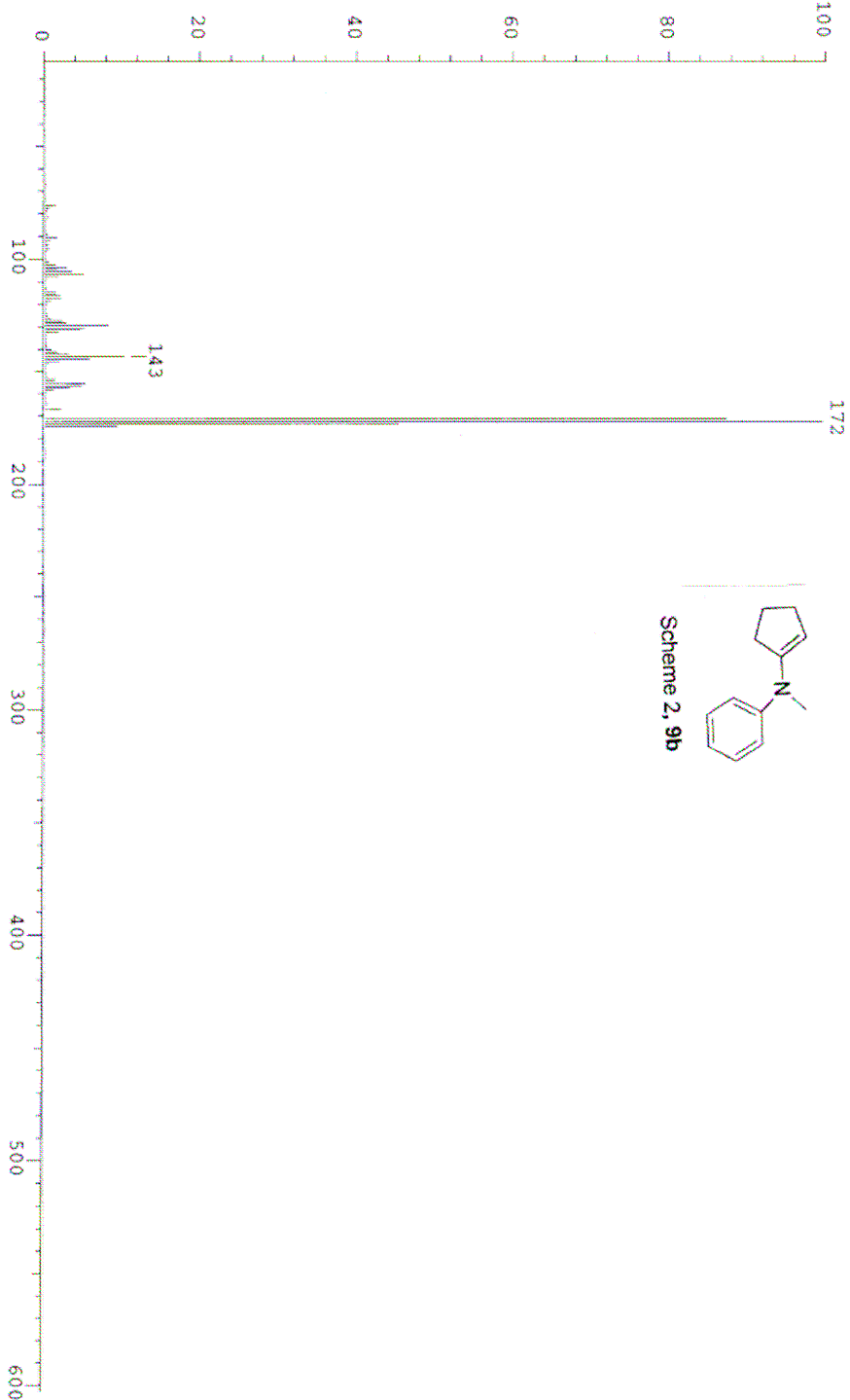




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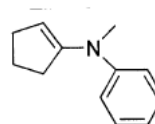


Scheme 2, 9b

Date: Thu Jun 2 14:53:55 2005 ICIS: 8.3.0 SP2 for OSF1 (V4.0) build 98-238 from 26-Aug-98

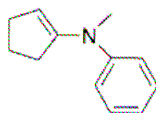
LIST: fin031694 02-JUN-05 Elapse: 00:28.6 15
 Samp: CV-1180 Start : 14:52:49 18
 Comm: 70 eV EI
 Mode: EI +Q1MS LMR UP LR Study : ms services
 Oper: kh Client: Venkat Inlet :
 Base: 172.4 Inten : 475876 Masses: 15 > 600
 Norm: 172.4 RIC : 1693267 #peaks: 595
 Peak: 1000.00 mmu

No.	Mass	Intensity	%RA	%RIC	Flags
1	129.	39149	8.23	2.31	F
2	143.	48766	10.25	2.88	F
3	171.	417179	87.67	24.64	F
4	172.	475876	100.00	28.10	F
5	173.	216118	45.41	12.76	F
6	175.	44273	9.30	2.61	F



Scheme 2, **9b**

Manual Peak Matching Report For Accurate Mass Determination



Scheme 2, 9b

Theoretical mass	Experimental mass	PFK matching mass	Deviation*
173.12045	173.12069	168.98882	1.4 ppm

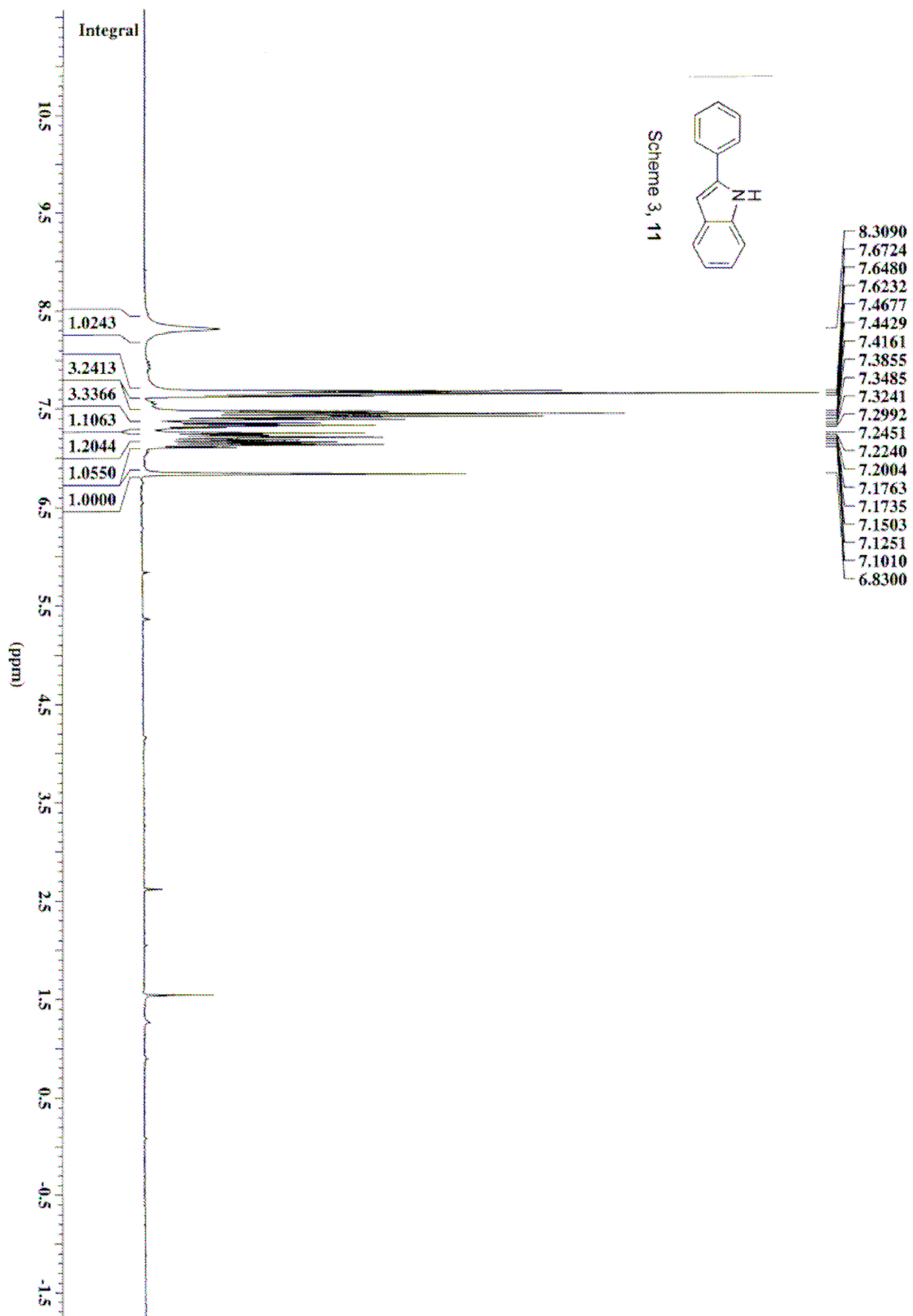
* The deviation is obtained from the following equation:

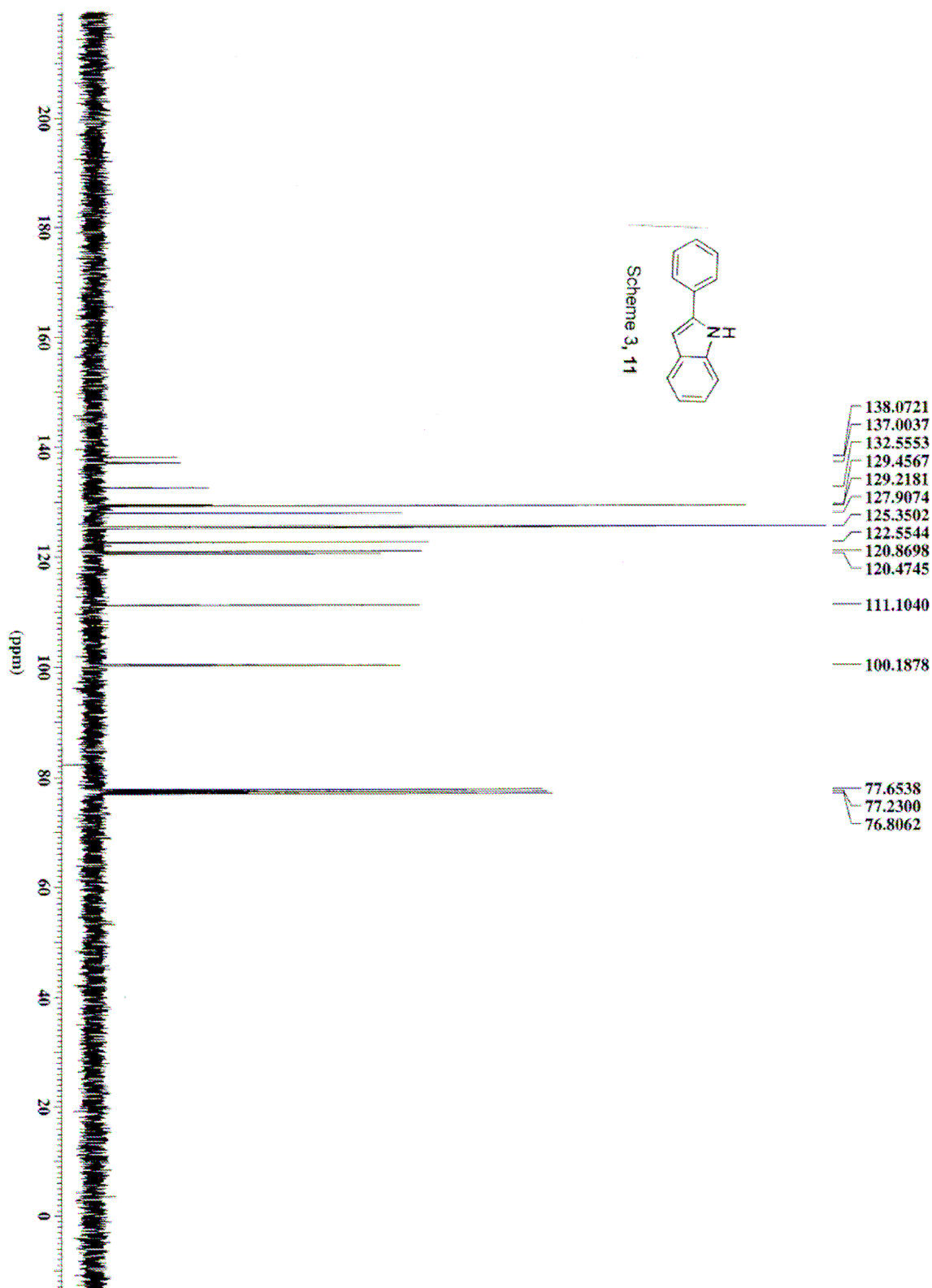
$$\text{deviation} = \frac{\text{experimental mass} - \text{theoretical mass}}{\text{nominal mass}}$$

Where nominal mass takes in account only ^{12}C , ^1H , ^{16}O , ^{14}N etc...

Theoretical mass correspond to the mass of the most abundant isotope peak

12





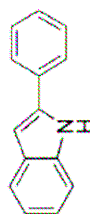
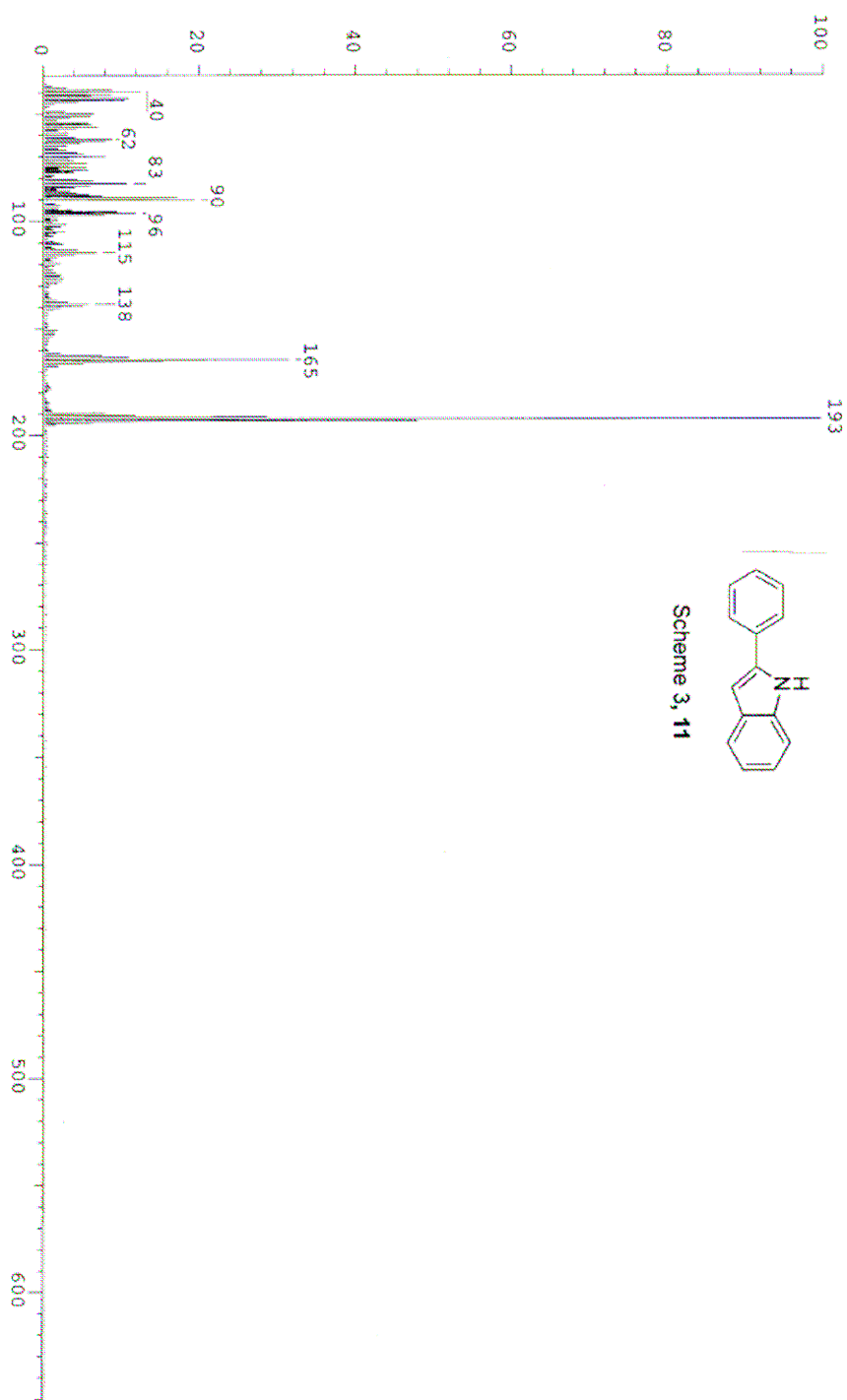
SPEC: fin031846.dat (15-JUN-05 14:35:40)
Samp: CVR-1269
Comm: 70 ev EI
Oper: kh
Base: 192.68
Peak: 1000.0 mmu
Scan 10 @ 0.36 min (EI +QIMS LMR UP LR)

Study: MS services
Masses: 35.01 > 650.00
Intensity: 162146

Scans: 1 > 11

Client: Venkate
#peaks: 583
RIC: 1163942

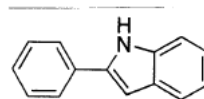
1.6E+05



Scheme 3, 11

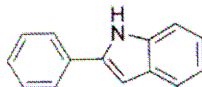
LIST: fin031846 15-JUN-05 Elapse: 00:21.6 10
 Samp: CVR-1269 Start : 14:35:40 11
 Comm: 70 eV EI
 Mode: EI +Q1MS LMR UP LR Study : MS services
 Oper: kh Client: Venkate Inlet :
 Base: 192.7 Inten : 162146 Masses: 35 > 650
 Norm: 192.7 RIC : 1163942 #peaks: 583
 Peak: 1000.00 mmu

14593					
No.	Mass	Intensity	%RA	%RIC	Flags
1	40.	19960	12.31	1.71	F
2	43.	17593	10.85	1.51	FM
3	43.	16595	10.23	1.43	FM
4	83.	17228	10.62	1.48	FM
5	89.	27649	17.05	2.38	F
6	90.	31331	19.32	2.69	F
7	95.	15088	9.31	1.30	F
8	96.	19006	11.72	1.63	F
9	163.	17640	10.88	1.52	FM
10	165.	50875	31.38	4.37	F
11	166.	24732	15.25	2.12	FM
12	191.	18818	11.61	1.62	F
13	192.	46487	28.67	3.99	F
14	193.	162146	100.00	13.93	F
15	194.	77834	48.00	6.69	F



Scheme 3, 11

Manual Peak Matching Report For Accurate Mass Determination



Scheme 3, 11

Theoretical mass	Experimental mass	PFK matching mass	Deviation*
193.08915	193.08945	180.98882	1.5 ppm

* The deviation is obtained from the following equation:

$$\text{deviation} = \frac{\text{experimental mass} - \text{theoretical mass}}{\text{nominal mass}}$$

Where nominal mass takes in account only ^{12}C , ^1H , ^{16}O , ^{14}N etc...

Theoretical mass correspond to the mass of the most abundant isotope peak

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